

Comprehensive Guide to Reaction Rates, Equilibrium, and Kinetics

Effect of Temperature on Reaction Rate

Increasing the temperature generally increases the rate of a chemical reaction by providing more energy to the reactant molecules. This results in more frequent and energetic collisions, which increases the likelihood of overcoming the activation energy barrier necessary for the reaction to proceed. As a consequence, higher temperatures typically lead to faster reactions, assuming other conditions remain constant.

Effect of Reactant Concentration on Reaction Rate

Increasing the concentration of reactant molecules typically accelerates the reaction rate. This is because a higher concentration means more particles are present in a given volume, increasing the frequency of collisions between reactant molecules. Since reactions occur when particles collide with sufficient energy and proper orientation, more collisions generally lead to a higher probability of successful reactions per unit time.

Role of Catalysts in Chemical Reactions

Catalysts speed up reactions by lowering the activation energy required for the reaction to proceed. They achieve this by providing an alternative reaction pathway. Importantly, catalysts are not consumed during the process; they remain unchanged at the end of the reaction. While catalysts increase the reaction rate, they do not alter the position of equilibrium; the equilibrium constant remains unaffected.

Guide Reaction Rates And Equilibrium

W. Christopher King



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