

READING COMPREHENSION

HS-PS1-2 NGSS ALIGNED

Chemical Bonding and Molecular Properties

5 Questions Type (25+ Questions)

➤ 10 Multiple Choice Questions

➤ 5 True or False Statements

➤ 5 Short Answer Questions

➤ 5 Fill in the Blanks Questions

➤ 2 Text Evidence Questions



**GOOGLE FORM
AND
PRINTABLE SHEET**

Answer Sheet Included

Name: _____

Topic: Chemical Bonding and Molecular Properties

Chemical Bonding and Molecular Properties

Chemical bonding describes the forces that hold atoms together in compounds. There are two main types of bonds: ionic and covalent. Ionic bonds form when electrons transfer from one atom to another, creating charged ions that attract each other. Covalent bonds occur when atoms share electrons to fill their outer shells, forming molecules.

Molecules formed by covalent bonds can be polar or nonpolar. Polarity depends on the difference in electronegativity between atoms. Electronegativity measures an atom's ability to attract electrons in a bond. When two atoms with different electronegativities bond, electrons are shared unequally, causing a polar molecule with positive and negative ends. Nonpolar molecules share electrons evenly and do not have charged poles.

Molecular properties such as melting point, boiling point, and solubility depend on the types of bonding and molecular shape. Polar molecules often dissolve in water because water is polar too, while nonpolar molecules dissolve better in nonpolar solvents. Understanding chemical bonding and molecular polarity helps chemists predict how substances interact and behave in different environments.

01. Mark the best answer for each question.

1. What type of bond involves the transfer of electrons?

- a. Ionic bond
- b. Covalent bond
- c. Hydrogen bond
- d. Metallic bond

2. How do atoms share electrons in a covalent bond?

- a. By transferring electrons completely
- b. By sharing neutrons
- c. By exchanging protons
- d. By sharing electrons to fill outer shells

3. What determines if a molecule is polar?

- a. Difference in electronegativity between atoms
- b. Number of atoms
- c. Molecular mass
- d. Temperature

4. What does electronegativity measure?

- a. The number of electrons in an atom
- b. The number of protons
- c. An atom's ability to attract electrons in a bond
- d. The mass of an atom

5. What kind of molecule shares electrons evenly?

- a. Radical
- b. Ionic compound
- c. Nonpolar molecule
- d. Polar molecule

Grade 10

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Cambridge



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