

ANSWER KEY

Mixed Stoichiometry Problems



a). How many moles of H_2 would be required to produce 5.0 moles of water?

5.0 moles water

b). What mass of H_2O is formed when H_2 reacts with 384 g of O_2 ?

432g H_2O



a). Balance this equation

Look above.

b). What mass of H_2SO_4 would be required to react with 0.75 mol of NaOH ?

36.8 = 37g H_2SO_4

d). If 30.0 g of H_2SO_4 reacts with sodium hydroxide, what mass of water is produced?

11.0g NaOH



a). How many grams of SnF_2 are produced from the reaction of 30.0 g of HF ? 117g SnF_2

b) How many moles of HF will be needed to react with 48 grams of Tin, Sn?

0.81 moles HF

4. In a spacecraft, the carbon dioxide exhaled by astronauts can be removed by its reaction with lithium hydroxide, LiOH , according to the following chemical equation.



If 20.0 moles of CO_2 is exhaled, the average amount exhaled by a person each day, how much (in grams) $\text{Li}_2\text{CO}_3(\text{s})$ is produced?

148g Li_2CO_3



a). Balance the equation.

See above.

b). How many moles of sodium are required to produce 5.0 moles of NaOH , sodium hydroxide?

5.0 moles Na

c). How many moles of H_2 are produced in the reaction with 5.0 grams of water?

0.14 moles H_2

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KJ Lindholm-Leary



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