

## Unit 7 – Equilibrium (7-9%)

Some reactions can occur in both forward and reverse directions, sometimes proceeding in each direction simultaneously.

**7.1 Intro to Equilibrium:** Explain the relationship between the occurrence of a reversible chemical or physical process, and the establishment of equilibrium, to experimental observations.

- Many observable processes are reversible. Examples include evaporation and condensation of water, absorption and desorption of a gas, or dissolution and precipitation of a salt. Some important reversible chemical processes include the transfer of protons in acid-base reactions and the transfer of electrons in redox reactions.
- When equilibrium is reached, no observable changes occur in the system. Reactants and products are simultaneously present, and the concentrations or partial pressures of all species remain constant.
- The equilibrium state is dynamic. The forward and reverse processes continue to occur at equal rates, resulting in no net observable change.
- Graphs of concentration, partial pressure, or rate of reaction versus time for simple chemical reactions can be used to understand the establishment of chemical equilibrium.

**7.2 Direction of Reversible Reactions:** Explain the relationship between the direction in which a reversible reaction proceeds and the relative rates of the forward and reverse reactions.

# Unit 7 Objectives

## Equilibrium

A system at equilibrium depends on the relationships between concentrations, partial pressures of chemical species, and equilibrium constant  $K$ .

AP Chemistry

REFERENCE



# Modeling Chemistry Unit 7 Objectives

**Thomas Griffiths**



**Modeling Chemistry Unit 7 Objectives:**

## Reviewing **Modeling Chemistry Unit 7 Objectives**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Modeling Chemistry Unit 7 Objectives**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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