

Making Copper—A Stoichiometry Experiment

ANALYSIS

1. How many grams of copper(II) chloride dihydrate ($\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$) and aluminum were used in the experiment? Show calculations where relevant.

$$\begin{array}{l}
 \text{3 moles} \\
 3 (170.5 \text{ g}) \\
 = 511.5 \text{ g}
 \end{array}
 \quad
 \begin{array}{l}
 \text{2 moles} \\
 2 (27 \text{ g}) \\
 = 54 \text{ g}
 \end{array}
 \quad
 \begin{array}{l}
 \text{3 moles} \\
 3 (63.5) \\
 = 190.5 \text{ g}
 \end{array}$$

Answer: 511.5 g

Answer: 54 g

2. Balance the equation for the reaction between $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ and Al.



3. Use the masses of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ and Al from question #1 to calculate the mass of Cu (theoretical yield) produced. (Note: When finding the molar mass of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ be sure to add the mass of 2 moles of water to the molar mass of CuCl_2).

$$\frac{190.5}{511.5} \times 100 = 37.24\%$$

Answer: 37.24%

Fe And Copper Chloride Dihydrate Stoichiometry

Jianjun Gao



Fe And Copper Chloride Dihydrate Stoichiometry:

Fe And Copper Chloride Dihydrate Stoichiometry Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the energy of words has been evident than ever. They have the capacity to inspire, provoke, and ignite change. Such is the essence of the book **Fe And Copper Chloride Dihydrate Stoichiometry**, a literary masterpiece that delves deep into the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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