

Chem 162A

Chapter 14. Covalent Bonding: Orbitals

Hybridization and the LE Model of Bonding

- Lewis structures of molecules
- prediction of geometry of molecules
- hybrid orbitals (sp^3 , sp^2 , sp , dsp^2 , d^2sp^3)
- interpretation of structure and bonding

Molecular Orbital Model of Bonding in Molecules

- molecular orbital diagrams
- bond order
- magnetism

Molecular Spectroscopy

- Electronic Spectroscopy
- Vibrational/Rotational Spectroscopies
- Nuclear Magnetic Resonance (NMR) Spectroscopy

14.1 Hybridization and the LE Model of Bonding

- Atomic bonding involves only valence orbitals
- Methane, CH_4



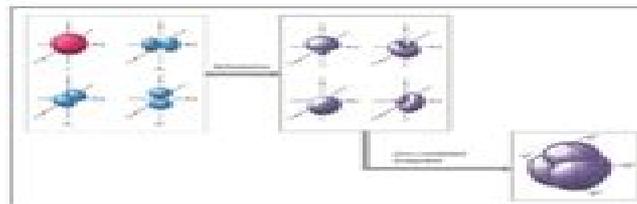
- H atoms in CH_4 will use $1s$ orbitals
- Of the two types of orbitals ($2s$ and $2p$) which will C atoms use for bonding in CH_4 ?
 - If both are used: 2 different types of C-H bonds (Contrary to experimental facts)
 - Neither of the "native" atomic orbitals of C atoms are used; instead, new hybrid orbitals are used.

Hybridization of atomic orbitals

- The mixing of the "native" atomic orbitals to form special orbitals for bonding is called **hybridization**.
- The 4 new **equivalent** orbitals formed by mixing the one $2s$ and three $2p$ orbitals are called sp^3 orbitals.
- The carbon atom is said to undergo sp^3 hybridization, i.e. is sp^3 hybridized.
- Energy-level diagram showing the sp^3 hybridization.



Figure 14.3: native $2s$ and three $2p$ atomic orbitals characteristic of a free carbon atom are combined to form a new set of four sp^3 orbitals.



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Table of Contents General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide

1. Understanding the eBook General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - The Rise of Digital Reading General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Personalized Recommendations
 - General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide User Reviews and Ratings
 - General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide and Bestseller Lists
5. Accessing General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide Free and Paid eBooks
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- 6. Navigating General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide Compatibility with Devices
 - General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Highlighting and Note-Taking General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Interactive Elements General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
- 8. Staying Engaged with General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
- 9. Balancing eBooks and Physical Books General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Setting Reading Goals General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Fact-Checking eBook Content of General Chemistry 2 Chapter Covalent Bonding Orbitals Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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