

# Final study guide question and answers

What kind of equilibrium exists and under what conditions? **Chemical equilibrium**  
Equilibrium is a dynamic state where the forward and reverse reactions occur at equal rates, resulting in no net change in the concentrations of reactants and products.

What happens if you change the concentration of a reactant? **Le Chatelier's Principle**  
If the concentration of a reactant is increased, the system will shift to the right to counteract the change and reach a new equilibrium.

What happens if you change the temperature? **Le Chatelier's Principle**  
If the temperature is increased, the system will shift in the direction that absorbs heat (endothermic reaction).

What happens if you change the pressure? **Le Chatelier's Principle**

If the pressure is increased, the system will shift towards the side with fewer moles of gas.

What happens if you change the volume? **Le Chatelier's Principle**

If the volume is decreased, the pressure is increased, and the system will shift towards the side with fewer moles of gas.

What happens if you change the catalyst? **Le Chatelier's Principle**

A catalyst speeds up both the forward and reverse reactions equally, so it does not affect the position of equilibrium.

What happens if you change the initial concentrations? **Le Chatelier's Principle**  
The system will shift to reach a new equilibrium based on the initial concentrations.

What happens if you change the initial pressure? **Le Chatelier's Principle**

If the initial pressure is increased, the system will shift towards the side with fewer moles of gas.

What happens if you change the initial volume? **Le Chatelier's Principle**

If the initial volume is decreased, the system will shift towards the side with fewer moles of gas.

What are some examples of chemical equilibrium? **Chemical Equilibrium**  
Examples include the Haber process for ammonia synthesis, the dissociation of acetic acid, and the reaction between nitrogen and oxygen to form nitric oxide.

# Nate Quick Study Guide

**JE Gale**



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