

# Final study guide question and answers

What kind of equilibrium exists and under what conditions? **Chemical equilibrium** is a dynamic equilibrium between two or more states of matter, where the concentrations of the reactants and products are constant over time.

What happens if you change the conditions? **Le Chatelier's principle** states that if a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will adjust itself to counteract the change and restore a new equilibrium.

What happens if you change the volume of the container? **Changing the volume** of a container at constant temperature and pressure will shift the equilibrium position. If the volume is increased, the equilibrium will shift towards the side with more moles of gas. If the volume is decreased, the equilibrium will shift towards the side with fewer moles of gas.

What happens if you change the temperature? **Changing the temperature** of a system at constant pressure and volume will shift the equilibrium position. For an exothermic reaction, increasing the temperature will shift the equilibrium towards the reactants. For an endothermic reaction, increasing the temperature will shift the equilibrium towards the products.

What happens if you change the concentration of the reactants or products? **Changing the concentration** of a reactant or product will shift the equilibrium position. Increasing the concentration of a reactant will shift the equilibrium towards the products, and increasing the concentration of a product will shift the equilibrium towards the reactants.

What happens if you change the pressure? **Changing the pressure** of a system at constant temperature and volume will shift the equilibrium position. Increasing the pressure will shift the equilibrium towards the side with fewer moles of gas, and decreasing the pressure will shift the equilibrium towards the side with more moles of gas.

What happens if you change the catalyst? **A catalyst** does not change the equilibrium position of a reaction. It only speeds up the rate at which equilibrium is reached.

What happens if you change the initial concentrations? **Changing the initial concentrations** of the reactants or products will shift the equilibrium position. Increasing the initial concentration of a reactant will shift the equilibrium towards the products, and increasing the initial concentration of a product will shift the equilibrium towards the reactants.

What happens if you change the initial volume? **Changing the initial volume** of a container at constant temperature and pressure will shift the equilibrium position. If the initial volume is increased, the equilibrium will shift towards the side with more moles of gas. If the initial volume is decreased, the equilibrium will shift towards the side with fewer moles of gas.

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What happens if you change the initial pressure? **Changing the initial pressure** of a system at constant temperature and volume will shift the equilibrium position. Increasing the initial pressure will shift the equilibrium towards the side with fewer moles of gas, and decreasing the initial pressure will shift the equilibrium towards the side with more moles of gas.

What happens if you change the initial concentration of the reactants or products? **Changing the initial concentration** of a reactant or product will shift the equilibrium position. Increasing the initial concentration of a reactant will shift the equilibrium towards the products, and increasing the initial concentration of a product will shift the equilibrium towards the reactants.

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What happens if you change the initial temperature of the system? **Changing the initial temperature** of a system at constant pressure and volume will shift the equilibrium position. For an exothermic reaction, increasing the initial temperature will shift the equilibrium towards the reactants. For an endothermic reaction, increasing the initial temperature will shift the equilibrium towards the products.

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**Spencer A. Rathus**



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