

9.1 Assessment

3. The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8.
4. from the number of electrons lost.
5. -ite or -ate
6. Group 1A metals, 1+; Group 3A (aluminum), 3+; Group 5A nonmetals, 3-
7. A monoatomic anion is a single atom with a negative charge; a polyatomic anion is two or more bound atoms with a negative charge.
8. a.) K^+ , cation, potassium ion
b.) O^{2-} anion, oxide ion
c.) Sn^{2+} cation, tin (II)
d.) Br^- anion, bromide ion
e.) Be^{2+} cation beryllium ion
f.) Co^{3+} cation, cobalt(III) ion
9. a.) NH_4^+
b.) Sn^{2+}
c.) CrO_4^{2-}
d.) NO_3^-

9.2 Assessment

14. Write the name of the cation followed by the name of the anion.
15. Write the symbol of the cation followed by the symbol of the anion, then use subscripts to balance the charges.
16. Write the symbol for the metal ion followed by the formula of the polyatomic ion and balance the charges. Name the cation first followed by the anion.
17. a.) $BeCl_2$
b.) Cs_2S
c.) NaI
d.) SrO
18. a.) $Cr(NO_2)_3$
b.) $NaClO_4$
c.) $Mg(HCO_3)_2$
d.) $Ca(C_2H_3O_2)_2$
19. a.) incorrect; charges are not balanced $MgSO_4$
b.)
c.) incorrect; charges are not balanced $BeCl_2$

9.3 Assessment

20. Prefixes indicate the number of atoms of each element in a molecule of the compound.

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