

Worksheet # 10 - Stoichiometry

1) You have place orders for supplies Togo's

a) They make 125 sandwiches a day. Assume that they put three slices of tomato on each sandwich and that one tomato provides six slices. How many sandwiches can they make if all you have left is 145 tomatoes? Should you get tomatoes delivered today or can you wait until tomorrow to have them delivered?

b) A meatball sandwich is made from 2 slices of cheese, three meatballs, one bun and three slices of tomato. Suppose you have 345 buns, 567 meatballs, 678 slices of tomato and 456 slices of cheese, how many meatball sandwiches can you make? How many slices of tomato were leftover?

2) Methanol, CH_3OH , is a possible automobile fuel. Energy is released when it burns in the presence of oxygen. (Balance the reaction first!)



a) One pound of methanol (454g) is reacted with excess oxygen. How many grams of CO_2 do you make?

b) Now suppose you have 454g of methanol and 454 g of O_2 , which reactant will you run out of first? How much H_2O will you make?

3) Sand, SiO_2 , is resistant to all acids except HF. (Balance the rxn first!) This also happens to be the reaction responsible for the etching of Si wafers with HF. The native oxide layer on the wafer is removed by dunking the wafer in HF.



a) How many grams of HF were required to produce 345 g of SiF_4 ?

b) Suppose you combine 576 g of SiO_2 and 345 g of HF. How many grams of H_2O do you produce?

4) Two common antacids are milk of magnesia, composed predominantly of $\text{Mg}(\text{OH})_2$, and Amphogel (a cousin of Maalox) which contains $\text{Al}(\text{OH})_3$.

a) Given equal number of moles of each compound, which one neutralizes the larger amount of stomach acid?

b) Given equal masses of each compound, which one neutralizes the larger amount of stomach acid?

(The trick here is to just pick an amount and try out each problem. You need balanced molecular reactions showing how HCl reacts with each of the hydroxides above.)

c) Tums contains CaCO_3 . We balance the reaction for this compound with HCl in class. Given equal masses of $\text{Al}(\text{OH})_3$ and CaCO_3 , which one will neutralize the most HCl?

5) The combustion of octane, C_8H_{18} , proceeds as listed below. Suppose you combine 3.55 g of octane with 678 g of O_2 .



a) How many grams of CO_2 can be produced?

b) How many grams of unreacted reactant are left over once the limiting reactant is all gone?

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Arturo Cuomo



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