

PROBLEM 14.2

KNOWN: Partial pressures and temperature for a mixture of CO_2 and N_2 .

FIND: Molar concentration, mass density, mole fraction and mass fraction of each species.

SCHEMATIC:

$$P_A = P_B = 1 \text{ bar}$$

$$T = 298 \text{ K}$$

$$A \rightarrow \text{CO}_2, \rho_A = 44 \text{ kg/kmol}$$

$$B \rightarrow \text{N}_2, \rho_B = 28 \text{ kg/kmol}$$

ASSUMPTIONS: (1) Perfect gas behavior.

ANALYSIS: From the equation of state for an ideal gas,

$$C_i = \frac{P_i}{RT}$$

Hence, with $P_A = P_B$,

$$C_A = C_B = \frac{1 \text{ bar}}{8.314 \times 10^{-2} \text{ m}^3 \cdot \text{bar/kmol} \cdot \text{K} \times 298 \text{ K}}$$

$$C_A = C_B = 0.040 \text{ kmol/m}^3$$

With $\rho_i = \rho_i C_i$, it follows that

$$\rho_A = 44 \text{ kg/kmol} \times 0.04 \text{ kmol/m}^3 = 1.76 \text{ kg/m}^3$$

$$\rho_B = 28 \text{ kg/kmol} \times 0.04 \text{ kmol/m}^3 = 1.13 \text{ kg/m}^3$$

Also, with

$$x_i = C_i / \sum_j C_j$$

find

$$x_A = x_B = 0.04 / 0.08 = 0.5$$

and with

$$m_i = \rho_i / \sum_j \rho_j$$

find

$$m_A = 1.76 / (1.76 + 1.13) = 0.61$$

$$m_B = 1.13 / (1.76 + 1.13) = 0.39$$

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William S. Janna



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