

Ionic Chemical Formulas/PreAICE Chemistry

NAME: _____ PERIOD _____

Directions: Using the Chemical Formulas program sheet, fill in the table below. Use subscripts & superscripts. Remember to use roman numerals to show the valence of the transition metals in the compound name.

METAL CATION (+)	CATION FORMULA	NONMETAL ANION (-)	ANION FORMULA	COMPOUND FORMULA	COMPOUND NAME
sodium	Na ⁺¹	hydroxide	OH ⁻	NaOH	sodium hydroxide
magnesium	Mg ⁺²	phosphate	PO ₄ ⁻³	Mg ₃ (PO ₄) ₂	magnesium phosphate
calcium	Ca ⁺²	chloride	Cl ⁻	CaCl ₂	calcium chloride
gold(I)	Au ⁺¹	sulfite	SO ₃ ⁻²	Au ₂ SO ₃	gold(I) sulfite
copper(II)	Cu ⁺²	hydroxide	OH ⁻	Cu(OH) ₂	copper(II) hydroxide
Iron(III)	Fe ⁺³	telluride	Te ⁻²	Fe ₂ Te ₃	iron(III) telluride
magnesium	Mg ⁺²	bromide	Br ⁻¹	MgBr ₂	magnesium bromide
manganese(II)	Mn ⁺²	sulfide	S ⁻²	MnS	manganese(II) sulfide
boron	B ⁺³	oxide	O ⁻²	B ₂ O ₃	boron oxide
potassium	K ⁺	nitrate	NO ₃ ⁻	KNO ₃	potassium nitrate
barium	Ba ⁺²	bromide	Br ⁻	BaBr ₂	barium bromide
strontium	Sr ⁺²	sulfate	SO ₄ ⁻²	SrSO ₄	strontium sulfate
iron(III)	Fe ⁺³	oxide	O ⁻²	Fe ₂ O ₃	iron(III) oxide
silver(I)	Ag ⁺¹	chloride	Cl ⁻	AgCl	silver(I) chloride
aluminum	Al ⁺³	carbonate	CO ₃ ⁻²	Al ₂ (CO ₃) ₃	aluminum carbonate
gold(II)	Au ⁺²	nitride	N ⁻³	Au ₃ N ₂	gold(II) nitride
cesium	Cs ⁺¹	phosphate	PO ₄ ⁻³	Cs ₃ PO ₄	cesium phosphate
ammonium	NH ₄ ⁺	sulfide	S ⁻²	(NH ₄) ₂ S	ammonium sulfide
zinc(II)	Zn ⁺²	nitrite	NO ₂ ⁻	Zn(NO ₂) ₂	zinc(II) nitrite
chromium(I)	Cr ⁺¹	nitride	N ⁻³	Cr ₃ N	chromium(I) nitride
mercury(II)	Hg ⁺²	oxide	O ⁻²	HgO	mercury(II) oxide
sodium	Na ⁺	sulfite	SO ₃ ⁻²	Na ₂ SO ₃	sodium sulfite
magnesium	Mg ⁺²	oxide	O ⁻²	MgO	magnesium oxide

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