

## Ch. 12 Review

E-book p.414

14. The intermolecular forces between the particles determine the state of a substance. In a solid, the intermolecular forces are very strong and hold the particles together. In a liquid, the intermolecular forces are weaker and in a gas, the particles no longer experience intermolecular forces.

15. Intermolecular forces occur between particles and intramolecular forces occur within particles.

16. Hydrogen bonds: d. HF      Dispersion forces: a. H<sub>2</sub>

*Reasoning: The H<sub>2</sub> molecule is nonpolar therefore it can only have intermolecular dispersion forces. HF is a polar molecule (permanent dipole) with hydrogen bound to fluorine which means the intermolecular forces are due to hydrogen bonding.*

P. 424

18. In solids the particles are closer together than in liquids because of intermolecular attractions.

20. Soaps and detergents decrease the surface tension of water by breaking the hydrogen bonds, which allows the dirt to be carried away by the water.

23. A meniscus forms because the adhesive forces between water molecules and the glass are greater than the cohesive forces between the water, so the water rises along the inside of the cylinder walls.

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35. Kinetic energy of particles is directly proportional to their temperature.

51. A temporary dipole forms when one molecule is close to another molecule and their electron clouds repel each other. When the electron cloud shifts, it creates a greater electron density in one part of the molecule.

Permanent dipoles are found in polar molecules where one part of the molecule is always partially positive and the other part is always partially negative.

52. Dispersion forces are weaker because they form between temporary dipoles. Dipole-dipole forces are between permanent dipoles.

53. A hydrogen bond involves a large difference in electronegativity between hydrogen and a F, O, or N atom which makes the bond extremely polar and stronger.

# Modern Chemistry Ch 12 Review Answer Key

**M Carnoy**



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