

Molarity Worksheet

Name: _____

1. What does molarity mean?

$$\frac{\text{Number of moles of solute}}{\text{1 liter solution}}$$

2. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution?

$$\frac{4.53 \text{ mol LiNO}_3}{2.85 \text{ L soln}} = 1.59 \text{ M LiNO}_3$$

3. What is the molarity of a solution that contains 0.00372 moles hydrochloric acid in 2.39×10^{-2} liters of solution?

$$\frac{0.00372 \text{ mol HCl}}{2.39 \times 10^{-2} \text{ L soln}} = 0.156 \text{ M HCl}$$

4. A flask contains 85.5 g $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ (sucrose) in 1.00 liters of solution. What is the molarity?

$$\frac{85.5 \text{ g sucrose} \times 1 \text{ mol sucrose}}{1.00 \text{ L soln} \quad 342.34 \text{ g sucrose}} = 0.250 \text{ M sucrose}$$

5. A beaker contains 214.2 grams osmium (III) fluoride in 0.0673 liters of solution. What is the molarity?

$$\frac{214.2 \text{ g OsF}_3 \times 1 \text{ mol OsF}_3}{0.0673 \text{ L soln} \quad 247.23 \text{ g OsF}_3} = 12.9 \text{ M OsF}_3$$

6. Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.

$$\frac{1.54 \text{ mol K}_2\text{SO}_4}{0.125 \text{ L soln}} = 12.3 \text{ M K}_2\text{SO}_4$$

7. A beaker contains 36.45 grams ammonium chlorite in 2.36 liters of solution - calculate the molarity.

$$\frac{36.45 \text{ g NH}_4\text{ClO}_2 \times 1 \text{ mol NH}_4\text{ClO}_2}{2.36 \text{ L soln} \quad 85.50 \text{ g NH}_4\text{ClO}_2} = 0.181 \text{ M NH}_4\text{ClO}_2$$

8. What is the molarity of a solution that contains 14.92 grams magnesium oxalate in 3.65 ml of solution?

$$\frac{14.92 \text{ g MgC}_2\text{O}_4 \times 1 \text{ mol MgC}_2\text{O}_4}{0.00365 \text{ L soln} \quad 112.32 \text{ g MgC}_2\text{O}_4} = 36.4 \text{ mol MgC}_2\text{O}_4$$

Molarity Practice Unit 10 Lesson Key

Victor M. Corman



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Table of Contents Molarity Practice Unit 10 Lesson Key

1. Understanding the eBook Molarity Practice Unit 10 Lesson Key
 - The Rise of Digital Reading Molarity Practice Unit 10 Lesson Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Practice Unit 10 Lesson Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Practice Unit 10 Lesson Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Practice Unit 10 Lesson Key
 - Personalized Recommendations
 - Molarity Practice Unit 10 Lesson Key User Reviews and Ratings
 - Molarity Practice Unit 10 Lesson Key and Bestseller Lists

5. Accessing Molarity Practice Unit 10 Lesson Key Free and Paid eBooks
 - Molarity Practice Unit 10 Lesson Key Public Domain eBooks
 - Molarity Practice Unit 10 Lesson Key eBook Subscription Services
 - Molarity Practice Unit 10 Lesson Key Budget-Friendly Options
6. Navigating Molarity Practice Unit 10 Lesson Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Practice Unit 10 Lesson Key Compatibility with Devices
 - Molarity Practice Unit 10 Lesson Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Practice Unit 10 Lesson Key
 - Highlighting and Note-Taking Molarity Practice Unit 10 Lesson Key
 - Interactive Elements Molarity Practice Unit 10 Lesson Key
8. Staying Engaged with Molarity Practice Unit 10 Lesson Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Practice Unit 10 Lesson Key
9. Balancing eBooks and Physical Books Molarity Practice Unit 10 Lesson Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Practice Unit 10 Lesson Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity Practice Unit 10 Lesson Key
 - Setting Reading Goals Molarity Practice Unit 10 Lesson Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Practice Unit 10 Lesson Key
 - Fact-Checking eBook Content of Molarity Practice Unit 10 Lesson Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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