

- 5) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \times 329 \times .0821}{.09 \text{ atm}} = 2310.9 \text{ L}$$

- 7) If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V}$$

$$\frac{17 \cdot .0821 \cdot 340}{88.89 \text{ L}} = 5.34 \text{ atm}$$

- 8) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT}$$

$$\frac{.0821 \times 300}{.5 \text{ atm} \times 25 \text{ L}} = 1.97 \text{ mol}$$

- 9) If I have 21 moles of gas held at a pressure of 78 atm and a temperature of 900 K, what is the volume of the gas?

$$PV = nRT$$

$$V = \frac{nRT}{P}$$

$$\frac{21 \cdot 0.0821 \cdot 900}{78} = 19.89 \text{ L}$$

- 10) If I have 1.9 moles of gas held at a pressure of 5 atm and in a container with a volume of 50 liters, what is the temperature of the gas?

$$PV = nRT$$

$$T = \frac{PV}{nR}$$

$$\frac{5 \text{ atm} \cdot 50 \text{ L}}{1.9 \text{ mol} \cdot .0821} = 1602.67 \text{ K}$$

- 11) If I have 2.4 moles of gas held at a temperature of 97 °C and in a container with a volume of 45 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V}$$

$$\frac{2.4 \times .0821 \cdot 370}{45} =$$

1.62 atm

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**CO Houle**



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