

Gas Laws Practice

Equations and Constants

1 atm = 760 mm Hg

1 atm = 14.70 psi

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$PV = nRT$$

R = 0.0821 L·atm/mol·K

Avogadro's #: 6.02×10^{23}

Describing Gases

1. Standard atmospheric pressure is 1 atmosphere. Express standard atmospheric pressure in mm Hg, kilopascals, and bars.

The Gas Laws

2. A gas occupies a volume of 224 cm³ and a pressure of 1.51 bar. If the gas expands to a volume of 578 cm³, what will the new pressure be? (Assume the temperature remains constant.)
3. A cylinder with a constant volume of 2.80 L has a pressure of 32.0 psi at a temperature of 25.0 °C. If the cylinder is warmed to a temperature of 75.0 °C, what will be the pressure inside the cylinder?
4. A gas occupies a volume of 600.0 mL at a temperature of 25.00 °C. At what temperature would the gas occupy only half this volume?
5. In air-conditioning systems, compressed gases are allowed to expand, and this expansion results in cooling. A gas with a volume of 12.0 mL at a pressure of 8.0 bar at a temperature of 45 °C is allowed to expand to a volume of 40.0 mL at a pressure of 2.0 bar. What is the temperature of the gas after it expands?
6. What is the volume of 4.52 moles of gas, calculated at standard temperature and pressure?
7. What is the pressure of 12.5 moles of gas at a temperature of 360.0 K and a volume of 5.02 liters?
8. A helium balloon has a volume of 3.4 liters and a pressure of 1.05 atmospheres at 25 °C. How many moles of helium gas are in the balloon? How many grams of gas?

Diffusion and Effusion

9. What is the difference between diffusion and effusion?

Gas Laws Extra Practice

RD Boyd



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