

Chemistry Honors Practice Test: Kinetics/Equilibrium

Use the following statements for Q1 & Q2:

"A change in concentration can cause a change in equilibrium" (A)

True

SO THAT

"An affected system (previously at equilibrium) will shift to the direction that receives stress" (B)

X

False

1. Which of the following represents the validity of the statements?

- a. Both A and B are true
- b. Both A and B are not true
- c. Only A is true
- d. Only B is true

2. The same principle that dictates the above statements can also be used to

- Determine whether a reaction is exothermic or endothermic X
- Manipulate profits and yields toward a desired result ✓
- Identify the degree of equilibrium using β units
- Analyze the extent to which external factors influence equilibrium X
- Examine the spontaneity of a reaction with greater precision X

Le Chatelier's principle

3. The equilibrium constant, K_c , for the reaction $\text{H}_{2(g)} + \text{I}_{2(g)} \rightarrow 2\text{HI}_{(g)}$ is 55 at 425°C . The value for the equilibrium constant will be changed if

- a. Concentrations are given in atmospheres instead of mol/L.
- b. A catalyst is added to the reaction system.
- c. The size of a reaction vessel is doubled.
- d. The temperature is changed to 400°C . *only Temp can change K*
- e. Additional HI is added to the reaction system.

Use the equation below to answer Q4:



4. To increase the amount of product, one could

- Increase pressure *same mols of gas both sides*
- Increase temperature *would increase reactants*
- Add more Cl_2 *would increase reactants*
- d. None will work (a-c)
- All will work (a-c)

Kinetics And Equilibrium Practice Test Honors Chemistry

Iowa State University



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