

Study Card: Factors Affecting Reaction Rates

Reaction rates are a measure of how fast a concentration changes with respect to time. Several factors affect the rate of a reaction. These include reactant concentration, temperature, surface area of reactants, and the use of a catalyst.

Reactant Concentrations

Consider the reaction:

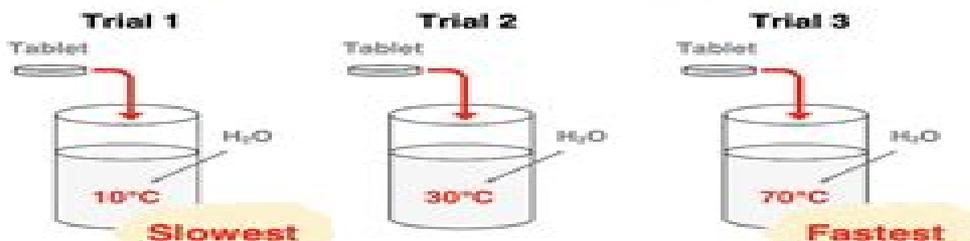


Vary the [HCl] in three trials and compare the results. Reaction rates increase with increasing reactant concentrations.



Temperature

An Alka Seltzer tablet dissolves in H_2O and reacts to produce CO_2 bubbles. Rates are easily observable. Vary the H_2O temperature for three trials and compare the results. Reaction rates increase with increasing temperature.



Surface Area (for Solids)

Chalk (CaCO_3) reacts with vinegar to produce CO_2 bubbles. Rates are easily observable. Run three trials, varying the surface area by breaking or grinding the chalk and compare the results. Reaction rates increase with increasing surface area.



Use of a Catalyst

Catalysts are substances that are designed to speed a reaction without being consumed. H_2O_2 decomposes slowly. But when a catalyst is added, the rate of decomposition can be stunningly rapid. The *elephant toothpaste demo* provides a classic example of this.



Fast

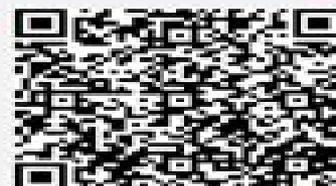
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