

## Reaction Quotient

How do you predict which direction a reaction will proceed to reach equilibrium?

### Why?

When a reaction reaches equilibrium there must be some non-negligible amount of every species in the reaction, otherwise the reaction cannot react in both directions. Knowing this, it is very easy to predict which direction a reaction will go to reach equilibrium when one of the components of the reaction has an initial concentration of zero. Many of the problems you have worked with thus far have some component at zero concentration, but real life does not work that way. Most of the time, the reaction in question has some measurable amount of every species. Deciding which way the reaction will go to reach equilibrium then becomes more challenging.

### Model 1 – A Theoretical Equilibrium

Trial 1	A (g) + B (g) $\leftrightarrow$ C (g)		
Initial	1.000 M	1.000 M	1.000 M
Change	$\leftarrow$		
Equilibrium	1.464 M	1.464 M	0.536 M

Trial 3	A (g) + B (g) $\leftrightarrow$ C (g)		
Initial	1.000 M	0.500 M	1.500 M
Change			
Equilibrium	1.864 M	1.364 M	0.636 M

Trial 5	A (g) + B (g) $\leftrightarrow$ C (g)		
Initial	1.400 M	1.200 M	0.400 M
Change			
Equilibrium	1.388 M	1.188 M	0.412 M

Trial 2	A (g) + B (g) $\leftrightarrow$ C (g)		
Initial	2.000 M	0.500 M	0.500 M
Change			
Equilibrium	2.150 M	0.650 M	0.350 M

Trial 4	A (g) + B (g) $\leftrightarrow$ C (g)		
Initial	1.600 M	1.000 M	0.400 M
Change			
Equilibrium	1.600 M	1.000 M	0.400 M

Trial 6	A (g) + B (g) $\leftrightarrow$ C (g)		
Initial	0.750 M	2.000 M	0.250 M
Change			
Equilibrium	0.675 M	1.925 M	0.325 M

1. Examine Model 1.
  - a. Write the theoretical chemical reaction that is used in the trials of Model 1.
  - b. If 0.50 M of reactant A reacts, predict the change in concentration of B and C.
2. What variables were changed in the different trials shown in Model 1?

# Equilibrium Pogil Activities For High School Chemistry Answers

**Jackie Acree Walsh**



## **Equilibrium Pogil Activities For High School Chemistry Answers:**

**POGIL** Shawn R. Simonson, 2023-07-03 Process Oriented Guided Inquiry Learning POGIL is a pedagogy that is based on research on how people learn and has been shown to lead to better student outcomes in many contexts and in a variety of academic disciplines Beyond facilitating students mastery of a discipline it promotes vital educational outcomes such as communication skills and critical thinking Its active international community of practitioners provides accessible educational development and support for anyone developing related courses Having started as a process developed by a group of chemistry professors focused on helping their students better grasp the concepts of general chemistry The POGIL Project has grown into a dynamic organization of committed instructors who help each other transform classrooms and improve student success develop curricular materials to assist this process conduct research expanding what is known about learning and teaching and provide professional development and collegiality from elementary teachers to college professors As a pedagogy it has been shown to be effective in a variety of content areas and at different educational levels This is an introduction to the process and the community Every POGIL classroom is different and is a reflection of the uniqueness of the particular context the institution department physical space student body and instructor but follows a common structure in which students work cooperatively in self managed small groups of three or four The group work is focused on activities that are carefully designed and scaffolded to enable students to develop important concepts or to deepen and refine their understanding of those ideas or concepts for themselves based entirely on data provided in class not on prior reading of the textbook or other introduction to the topic The learning environment is structured to support the development of process skills such as teamwork effective communication information processing problem solving and critical thinking The instructor s role is to facilitate the development of student concepts and process skills not to simply deliver content to the students The first part of this book introduces the theoretical and philosophical foundations of POGIL pedagogy and summarizes the literature demonstrating its efficacy The second part of the book focusses on implementing POGIL covering the formation and effective management of student teams offering guidance on the selection and writing of POGIL activities as well as on facilitation teaching large classes and assessment The book concludes with examples of implementation in STEM and non STEM disciplines as well as guidance on how to get started Appendices provide additional resources and information about The POGIL Project

**Questioning for Formative Feedback** Jackie Acree Walsh, 2022-05-20 When used effectively quality questions and student dialogue result in self regulated learners and formative feedback that reveals progress toward learning goals Learning knows no boundaries The potential for learning exists whenever and wherever we interact with our environment So how can we infuse school learning with the authenticity and excitement associated with real life experiences In Questioning for Formative Feedback Jackie Acree Walsh explores the relationship between questioning and feedback in K 12 classrooms and how dialogue serves as the bridge connecting the two Quality questioning productive dialogue and

authentic use of feedback are a powerful trifecta for addressing the needs of a new generation of learners. In fact, the skillful use of these three processes can fuel and accelerate the academic, social, and emotional learning of all students. In this book, Walsh provides a manual of practice for educators who want to engage students as partners in these processes. To that end, she offers the following features to help create a classroom in which everyone learns through intentional practice: Blueprints for coherent models of key processes and products; Tools and strategies to help you achieve identified outcomes; Protocols with step-by-step directions to complete an activity; Classroom artifacts of authentic classroom use, including links to 21 original videos produced exclusively for this book; Working together, questioning, dialogue, and feedback can transform learning for all. This book supports you in embracing and bringing that vision to fruition.

*Chemistry Equations and Answers* Mark Jackson, 2006-02-22. This 6-page study guide contains basic chemistry analysis and concepts designed specifically to aid science students.

*High School Chemistry* Vishal Mody, 2016-05-20. This chemistry booklet was created to help students specifically with the topic of Chemical Equilibrium in chemistry. This booklet has been made extremely concise yet explains the concepts in detail at the same time. Remember that this booklet is not designed to be your main study source but rather as an adjunct to your school teacher's notes. There are also lots of practice questions with detailed solutions at the end to solidify the concepts you have learned.

*Equilibrium Activity Diagrams* T. S. Bowers, K. J. Jackson, H. C. Helgeson, 2012-12-06. This book represents a revision and expansion of an earlier set of diagrams for temperatures from 25 to 300 °C along the equilibrium vapor-liquid curve for H<sub>2</sub>O. Helgeson, Brown, and Leeper (1969) summarized the activity diagrams in the following pages, which were generated over a six-year period from 1977 to 1983 in the Laboratory of Theoretical Geochemistry, otherwise known as Prediction Center, at the University of California, Berkeley. They represent the culmination of research efforts to generate a comprehensive and internally consistent set of thermodynamic data and equations for minerals, gases, and aqueous solutions at high pressures and temperatures. Among the many who contributed to the successful completion of this book, we are especially indebted to David Kirkham, John Walther, and George Flowers, who wrote program SUPCRT; Tom Brown, who created program DIAGRAM; and Eli Messinger, who generated the Tektronix plot routine to construct the diagrams. Ken Jackson and Terri Bowers both devoted an enormous amount of time and effort over the past six years to produce the diagrams in the following pages, some of which went through many stages of revision. Consequently, they appear as senior authors of this volume. It should be mentioned in this regard that their equal dedication to the project made it necessary to determine their order of authorship by flipping a coin.

**Equilibrium**, 1991 **Equilibrium Calculations** John T. Donoghue, 1971 **Solution Equilibria** F. R. Hartley, Christopher Burgess, R. M. Alcock, 1980

*Chemical Equilibrium* William Guenther, 1975-12. The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry, biochemistry, geochemistry, and interdisciplinary environmental fields. The pedagogical strategy I have adopted makes heavy use of detailed examples of

problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

**Equilibrium** Thomas R. Blackburn, 1969 [Chemical Equilibrium In a Nutshell](#) Brian Coleman, 2018-06-25

Understanding the math and minutiae of chemical equilibrium can be a tall task for anyone so why not enlist the help of a scientific squirrel to guide you on your journey Join Dr Wash as we dabble in equilibrium constants and other tools needed to predict chemical processes This book focuses on introductory concepts at the high school and early university level focusing on identifying equilibrium calculating K and Q discussing Le Chatelier's principle and tying equilibrium with the field of Thermodynamics Full of step by step instructions and practice questions this book aims to simplify one of the more complex topics found within the field of chemistry

*Principles of Chemical Equilibrium* Brian Craven, 1972

**Equilibrium Activity Diagrams** Teresa S. Bowers, Kenneth J. Jackson, Harold C. Helgeson, 1984 This book represents a revision and expansion of an earlier set of diagrams for temperatures from 25 to 300 C along the equilibrium vapor liquid curve for H<sub>2</sub>O Helgeson Brown 2 and Leeper 1969 The activity diagrams summarized in the following pages were generated over a six year period from 1977 to 1983 in the Laboratory of Theoretical Geochemistry otherwise known as Prediction Center at the University of California Berkeley They represent the culmination of research efforts to generate a comprehensive and internally consistent set of thermodynamic data and equations for minerals gases and aqueous solutions at high pressures and temperatures Among the many who contributed to the successful completion of this book we are especially indebted to David Kirkham John Walther and George Flowers who wrote program SUPCRT Tom Brown who created program DIAGRAM and Eli Messinger who generated the Tektronix plot routine to construct the diagrams Ken Jackson and Terri Bowers both devoted an enormous amount of time and effort over the past six years to produce the diagrams in the following pages some of which went through many stages of revision Consequently they appear as senior authors of this volume It should be mentioned in this regard that their equal dedication to the project made it necessary to determine their order of authorship by flipping a coin

**Equilibrium Activity Diagrams for Coexisting Minerals and Aqueous Solutions at Pressures and Temperatures to 5 KB and 600 O C** Teresa S. Bowers, 1984 **Chemical Equilibrium and Solutions** James E.

Banks,1967    **Equilibrium and Energetics** John Allen Berry-Richards,1972    **Equilibrium** William Edward  
Marsden,1996    Chemical Equilibrium William B. Guenther,1975    BASIC Equilibrium Calculations Cooper Harold  
Langford,1973    **Equilibrium III** ,1984

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