

## Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:



$$K = \frac{[\text{NH}_3]^2}{[\text{N}_2] [\text{H}_2]^3}$$



$$K = \frac{[\text{ICl}]^2}{[\text{Cl}_2]}$$



$$K = \frac{[\text{NO}] [\text{O}_2]^{1/2}}{[\text{NO}_2]}$$

2. The dissociation of acetic acid,  $\text{CH}_3\text{COOH}$ , has an equilibrium constant at  $25^\circ\text{C}$  of  $1.8 \times 10^{-5}$ . The reaction is



If the equilibrium concentration of  $\text{CH}_3\text{COOH}$  is 0.46 moles in 0.500 L of water and that of  $\text{CH}_3\text{COO}^-$  is  $8.1 \times 10^{-3}$  moles in the same 0.500 L, calculate  $[\text{H}^+]$  for the reaction.

$$K = \frac{[\text{CH}_3\text{COO}^-] [\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad \text{Plug in known values and solve.}$$

$$1.8 \times 10^{-5} = \frac{[8.1 \times 10^{-3} \text{ moles}/0.500 \text{ L}] [\text{H}^+]}{[0.46 \text{ moles}/0.500 \text{ L}]}$$

$$[\text{H}^+] = 1.0 \times 10^{-3} \text{ M}$$

3. Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

	catalyst	pressure	temperature	concentration
a. speed	presence increase	direct relationship	direct relationship	direct relationship
b. pos <sup>n</sup>	nothing	based on moles	based on enthalpy of rxn	inverse relationship

# Equilibrium Problems Practice

**Tobias Bleicker**



## **Equilibrium Problems Practice:**

This book delves into Equilibrium Problems Practice. Equilibrium Problems Practice is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Equilibrium Problems Practice, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Equilibrium Problems Practice
- Chapter 2: Essential Elements of Equilibrium Problems Practice
- Chapter 3: Equilibrium Problems Practice in Everyday Life
- Chapter 4: Equilibrium Problems Practice in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Equilibrium Problems Practice. This chapter will explore what Equilibrium Problems Practice is, why Equilibrium Problems Practice is vital, and how to effectively learn about Equilibrium Problems Practice.

3. In chapter 2, this book will delve into the foundational concepts of Equilibrium Problems Practice. The second chapter will elucidate the essential principles that need to be understood to grasp Equilibrium Problems Practice in its entirety.

4. In chapter 3, the author will examine the practical applications of Equilibrium Problems Practice in daily life. This chapter will showcase real-world examples of how Equilibrium Problems Practice can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Equilibrium Problems Practice in specific contexts. The fourth chapter will explore how Equilibrium Problems Practice is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Equilibrium Problems Practice. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Equilibrium Problems Practice.

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