

EXERCISES

MULTIPLE-CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of K_c .
- C. In an open system, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen monoxide, NO , is converted to nitrogen dioxide, NO_2 , occurs at equilibrium:



In a given closed system, the equilibrium constant for this reaction is 0.0010. If the initial concentration of NO at equilibrium is 0.0020 M, what is the equilibrium concentration of N_2O_2 at equilibrium?

- A. 0.0020 M
- B. 0.0010 M
- C. 0.0015 M
- D. 0.0005 M

10) Dissolved carbon dioxide, CO_2 , reacts with water, H_2O , to form, respectively, H_2CO_3 , H_2CO_3^* , H_2CO_3^* , species that react with H_2O to form bicarbonate ions, HCO_3^- , and hydroxonium ions, H_3O^+ , according to the equilibrium below:



11) Suppose that a solution of a substance reacts to produce an equilibrium with itself. At equilibrium, the forward reaction is irreversible. In this scenario, what change is made to the system, which one of the following changes is most likely to result in an increase in the concentration of H_3O^+ ions?

- A. Change the solution.
- B. Double the mass CO_2 gas.
- C. Add a few drops of phenolphthalein indicator.
- D. Add a few drops of 0.1 M sodium hydroxide, NaOH , solution.

Equilibrium Test 2 Multiple Choice Answers

Eric R. Dodge



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