

Essay Format

Introduction:

- **Hook:** Start with a compelling or thought-provoking sentence or question to grab the reader's attention.
- **Background Information:** Provide some context or background information related to the topic.
- **Thesis Statement:** Clearly state the main argument or purpose of the essay in one or two sentences.

Body Paragraphs:

- **Each paragraph should focus on a single point or idea that supports the thesis statement.**
- **Topic Sentences:** Start each paragraph with a clear topic sentence that introduces the main point of the paragraph.
- **Evidence and Examples:** Provide evidence, examples, statistics, or quotes to support your point. Be sure to cite your sources if necessary.
- **Explanation and Analysis:** Explain how the evidence or examples provided relate to the thesis and why they are significant.
- **Transition Sentences:** Use transition sentences to smoothly connect one paragraph to the next.

Conclusion:

- **Restate Thesis:** Begin the conclusion by restating the thesis statement, but do not simply repeat it word for word.
- **Summarize Main Points:** Summarize the main points discussed in the body paragraphs.
- **Broaden the Scope:** Connect the argument to a broader context or real-world implications.
- **Your Thoughts:** Offer some final thoughts, reflections, or suggestions related to the topic.
- **Closing Statement:** End with a strong closing statement that leaves a lasting impression on the reader.

ESSAY FORMAT TEMPLATE

Ideal for students and professionals, it simplifies essay structuring for impactful results.

2 Pages



Essay Word Document

Julia Schneider



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web elements in the second row of the periodic table place their electrons in the $2n$ shell as well as the $1n$ shell for instance lithium li has three electrons two fill the $1s$ orbital and the third is placed in the $2s$ orbital giving an electron configuration of $1s^2 2s^1$

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configuration $3s^2$ explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain Hund's rule

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are $He\ 2$ helium $1s^2$ $Ne\ 10$ neon $1s^2\ 2s^2\ 2p^6$

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