

Chapter Test A

Chapter: Chemical Equations and Reactions

Do the work provided below the letter of the item or place the item completion work underneath or next to where work appears.

1. The mass number of an atom is the sum of the number of protons and neutrons. Which of the following descriptions does not describe atoms that are chemically identical but not identical?
 - a. They have the same mass.
 - b. They have the same number of electrons.
 - c. They have the same number of protons and neutrons.
 - d. They all have the same chemical properties.
2. An element is described by chemical equations:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$
 - a. carbon hydrocarbon is produced.
 - b. carbon ions produced.
 - c. carbon is a reactant.
 - d. both a and c.
3. Which of the following relationships is always true for a balanced equation?
 - a. The products of a reaction will contain the same number of atoms of the elements.
 - b. All of the reactants will be used.
 - c. The number of molecules of each reactant and product is equal.
 - d. The number of molecules of each reactant and product will equal.
4. There is a negative enthalpy change for a chemical equation. Which of the following is true?
 - a. The reactants of the reaction are products.
 - b. A heat-releasing reaction will have been performed.
 - c. The reactants of the reaction are products.
 - d. All of the above.
5. The number of chemical equations that can be written for a reaction is
 - a. equal the number.
 - b. equal the number of products.
 - c. equal the number of reactants.
 - d. equal the coefficients.

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**Thirumalaisamy P.
Velavancorresponding**



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