

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Essentials

Standard Constant:

| Unit | Standard Pressure |
|-------|-------------------|
| atm | 1.013 |
| mmHg | 760 |
| cm Hg | 760 |
| torr | 760 |
| atm | 1.013 |

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

| | |
|-------------------|--|
| Boyle's Law: | $P_1V_1 = P_2V_2$ |
| Charles's Law: | $\frac{V_1}{T_1} = \frac{V_2}{T_2}$ |
| Gay-Lussac's Law: | $\frac{P_1}{T_1} = \frac{P_2}{T_2}$ |
| Avogadro's Law: | $\frac{V_1}{n_1} = \frac{V_2}{n_2}$ |
| Combined Gas Law: | $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$ |
| Graham's Law: | $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$ |

The constant R (used in the other gas laws) depends on the unit of Pressure:

- 1 atm, $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$
- 1 atm, $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$

Always Remember:

- Temperature must always be in Kelvin! ($K = ^\circ\text{C} + 273$)
- If you are given two different units of pressure, CONVERT!
- Create two columns before solving, initial and final, write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you're given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (MAYBE!)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things!

Charles' Law $V \propto T$

Directions: Complete each of the problems below. Show how you set up the problem and solve.

- The temperature inside my refrigerator is about 4°C. If I place a balloon in the refrigerator, will the temperature of the air inside the balloon increase or decrease? Explain your answer. Is this related to the temperature?
- A balloon is inflated to a volume of 1.0 L at 27°C. If the temperature is lowered to 0°C, what will the volume of the balloon be?

Gay-Lussac's Law $P \propto T$

- A gas is held at a constant volume of 1.0 L and a pressure of 1.0 atm. If the temperature is increased from 27°C to 127°C, what will the pressure be?
- A gas is held at a constant volume of 1.0 L and a pressure of 1.0 atm. If the temperature is decreased from 27°C to 0°C, what will the pressure be?

Boyle's Law $P \propto \frac{1}{V}$

- A gas is held at a constant temperature of 27°C and a pressure of 1.0 atm. If the volume is increased from 1.0 L to 2.0 L, what will the pressure be?
- A gas is held at a constant temperature of 27°C and a pressure of 1.0 atm. If the volume is decreased from 1.0 L to 0.5 L, what will the pressure be?

Avogadro's Law $V \propto n$

- A gas is held at a constant temperature and pressure of 27°C and 1.0 atm. If the volume is increased from 1.0 L to 2.0 L, how many moles of gas are added?
- A gas is held at a constant temperature and pressure of 27°C and 1.0 atm. If the volume is decreased from 1.0 L to 0.5 L, how many moles of gas are removed?

Combined Gas Law $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

- A gas is held at a constant temperature and pressure of 27°C and 1.0 atm. If the volume is increased from 1.0 L to 2.0 L, what will the temperature be?
- A gas is held at a constant temperature and pressure of 27°C and 1.0 atm. If the volume is decreased from 1.0 L to 0.5 L, what will the temperature be?

Graham's Law $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

- Hydrogen gas (H₂) effuses through a porous barrier 9.7 times faster than an unknown gas. What is the molecular mass of the unknown gas?
- Helium gas (He) effuses through a porous barrier 3.0 times faster than an unknown gas. What is the molecular mass of the unknown gas?

All Gas Laws Included!



Gas Law Problems Packet

Katrin Zwirgmaier



Gas Law Problems Packet:

Adopting the Beat of Appearance: An Emotional Symphony within **Gas Law Problems Packet**

In some sort of consumed by displays and the ceaseless chatter of instant conversation, the melodic splendor and emotional symphony developed by the written word frequently fade in to the backdrop, eclipsed by the relentless noise and interruptions that permeate our lives. Nevertheless, set within the pages of **Gas Law Problems Packet** a charming literary prize filled with natural feelings, lies an immersive symphony waiting to be embraced. Constructed by a masterful musician of language, that captivating masterpiece conducts visitors on a psychological trip, skillfully unraveling the concealed melodies and profound affect resonating within each cautiously crafted phrase. Within the depths of the emotional analysis, we shall examine the book is key harmonies, analyze its enthralling publishing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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