

13. Complete the following table for the three experiments in Model 2.

	Experiment C	Experiment D	Experiment E
Independent Variable	# of particles	Temp.	Internal External Pressure
Dependent Variable	Volume	Volume	Volume
Controlled Variable(s)	Temp / I+E pressure	External + Internal Pressure / # of particles	Temp / # of particles

14. Provide a molecular level explanation for the increase in volume among the balloons in Experiment C. (How often and/or how hard are the molecules hitting the sides of the container?)

of molecules \uparrow , so number of collisions \uparrow + Volume along with it

collisions \uparrow energy stays the same

15. Provide a molecular level explanation for the increase in volume among the balloons in Experiment D.

Temperature \uparrow increasing the kinetic energy of the particles as well
which \uparrow collisions thus \uparrow pressure

16. Provide a molecular level explanation for the increase in volume among the balloons in Experiment E.

if external forces decrease then it allows internal forces \uparrow
thus \uparrow volume

Gas Variables Pogil Answer

Claire Vallance



Gas Variables Pogil Answer:

Gases and Gas Laws Hobsons Publishing, PLC,1983-01-01 Gas Movements in the Long-period Variables C. E. R. Bruce,1957 Math in Gas and the Art of Linearization Ton van der Hoeven,2004 **An Introduction to the Gas Phase** Claire Vallance,2017-11-23 An Introduction to the Gas Phase is adapted from a set of lecture notes for a core first year lecture course in physical chemistry taught at the University of Oxford The book is intended to give a relatively concise introduction to the gas phase at a level suitable for any undergraduate scientist After defining the gas phase properties of gases such as temperature pressure and volume are discussed The relationships between these properties are explained at a molecular level and simple models are introduced that allow the various gas laws to be derived from first principles Finally the collisional behaviour of gases is used to explain a number of gas phase phenomena such as effusion diffusion and thermal conductivity

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