



GAS VOLUMES 1

$$\text{Volume (dm}^3\text{)} = 24 \times \text{moles}$$



The volume must be in dm³ (there are 1000 cm³ in 1 dm³).

$$\text{vol in dm}^3 = \frac{\text{vol in cm}^3}{1000}$$

Give all answers to 3 significant figures.

1 Find the volume of the following gases (measured at room temperature and pressure).

- a 4.00 moles of oxygen (O₂) **volume = 4.00 x 24 = 96.0 dm³**
- b 0.250 moles of methane (CH₄) **volume = 0.25 x 24 = 6.00 dm³**
- c 15.0 g of argon (Ar) **moles Ar = $\frac{15}{40} = 0.375 \text{ mol}$ volume = 0.375 x 24 = 9.00 dm³**
- d 0.220 g of carbon dioxide (CO₂) **moles CO₂ = $\frac{0.220}{44} = 0.00500 \text{ mol}$ volume = 0.00500 x 24 = 0.120 dm³**

2 Find the number of moles of the following gases (measured at room temperature and pressure).

- a 48.0 dm³ of carbon monoxide (CO) **moles CO = $\frac{48.0}{24} = 2.00 \text{ mol}$**
- b 1.20 dm³ of hydrogen (H₂) **moles H₂ = $\frac{1.20}{24} = 0.0500 \text{ mol}$**
- c 360 cm³ of oxygen (O₂) **moles O₂ = $\frac{360}{24000} = 0.01500 \text{ mol}$**

3 Find the mass of the following gases (measured at room temperature and pressure).

- a 7.20 dm³ of ammonia (NH₃) **moles NH₃ = $\frac{7.20}{24} = 0.300 \text{ mol}$ mass = 17 x 0.300 = 5.10 g**
- b 480 cm³ of nitrogen (N₂) **moles N₂ = $\frac{480}{24000} = 0.0200 \text{ mol}$ mass = 28 x 0.0200 = 0.560 g**
- c 100 cm³ of oxygen (O₂) **moles O₂ = $\frac{100}{24000} = 0.00417 \text{ mol}$ mass = 32 x 0.00417 = 0.133 g**

Gas Variable Model 1 Answer Key

Clifford Lane



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Table of Contents Gas Variable Model 1 Answer Key

1. Understanding the eBook Gas Variable Model 1 Answer Key
 - The Rise of Digital Reading Gas Variable Model 1 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Variable Model 1 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Variable Model 1 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Variable Model 1 Answer Key
 - Personalized Recommendations
 - Gas Variable Model 1 Answer Key User Reviews and Ratings
 - Gas Variable Model 1 Answer Key and Bestseller Lists
5. Accessing Gas Variable Model 1 Answer Key Free and Paid eBooks
 - Gas Variable Model 1 Answer Key Public Domain eBooks
 - Gas Variable Model 1 Answer Key eBook Subscription Services

- Gas Variable Model 1 Answer Key Budget-Friendly Options
- 6. Navigating Gas Variable Model 1 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Variable Model 1 Answer Key Compatibility with Devices
 - Gas Variable Model 1 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Variable Model 1 Answer Key
 - Highlighting and Note-Taking Gas Variable Model 1 Answer Key
 - Interactive Elements Gas Variable Model 1 Answer Key
- 8. Staying Engaged with Gas Variable Model 1 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Variable Model 1 Answer Key
- 9. Balancing eBooks and Physical Books Gas Variable Model 1 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Variable Model 1 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Variable Model 1 Answer Key
 - Setting Reading Goals Gas Variable Model 1 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Variable Model 1 Answer Key
 - Fact-Checking eBook Content of Gas Variable Model 1 Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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