

ACS GENERAL CHEMISTRY FINAL STUDY GUIDE

Avagadro's Number - $N_A = 6.023 \times 10^{23}$

Percent Yield - $(\text{Actual Yield}/\text{Theoretical Yield}) \times 100$

Percent Composition (of an Element) - $[(\text{moles of element} \times \text{molar mass of element})/(\text{molar mass of compound})] \times 100$

Dilution Equation - $M_1V_1 = M_2V_2$
Mass and Volume

Boyle's Law - $P_1V_1 = P_2V_2$
Pressure and Volume

Charles' Law - $V_1/T_1 = V_2/T_2$
Volume and Temperature

Ideal Gas Law - $PV = nRT$

Volume of an Ideal Gas at STP - 22.4 L

Density (Ideal Gas) - $(P \times \text{Molar Mass})/(RT)$

Definition of Entahipy - $H = E + PV$

Definition of Heat Capacity - $C = ms$
 $s = \text{specific heat}$

ion - different number of electrons

isotopes - different number of neutrons

Periodic Table Trends - As you go down or left atomic radius increases
As you go up or right ionization energy and electron affinity increases

Ionic Bonds - to charged pieces come together and then there are cations or anions

Covalent Bonds - are when two non metals share electrons

A Compound is probably soluble if it contains these anions: - Cl^-
 Br^-
 I^-

General Chemistry Final Study Guide

Mary McDonald



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