

How to Name Chemical Compounds

is it ionic (contains metal) or does it contain a polyatomic ion (not provided by H)?

otherwise...

Ionic compound

[cation's name] [anion's name]

main element	<ul style="list-style-type: none"> take element name! ex. Al^{3+}: aluminum if transition metal*, include charge in Roman numeral ex. Cr^{2+}: copper (II) 	<ul style="list-style-type: none"> get element's name, chop end, add "-ide" ex. F^{-}: fluorine → fluoride
poly-atomic	<ul style="list-style-type: none"> just use its name! ex. NH_4^{+}: ammonium 	<ul style="list-style-type: none"> just use its name! ex. CO_3^{2-}: carbonate

* exceptions (no Roman numerals):

silver ion: always Ag^{+}

zinc ion: always Zn^{2+}

cadmium ion: always Cd^{2+}

Molecular compound

is it $HClO_x$?

Acid!
root and numbers

[#]-[1st element name] [#]-[2nd element root]-"ide"

↑ skip if "mono"

↑ skip if 1st element is H

Numerical prefixes

- 1: mono-
- 2: di-
- 3: tri-
- 4: tetra-
- 5: penta-
- 6: hexa-
- 7: hepta-
- 8: octa-
- 9: nona-
- 10: deca-



Acids

Ignore all H⁺, and look at the name of the remaining anion.

Anion	Xide	Xate	Xate
Acid	hydro-X-ic acid	X-ic acid	X-ous acid
	ex. Br^{-} bromide HBr hydrobromic acid	ex. ClO_3^{-} chlorate $HClO_3$ chloric acid	ex. ClO_2^{-} chlorite $HClO_2$ chlorous acid
	ex. CO_3^{2-} carbonate H_2CO_3 carbonic acid	ex. PO_4^{3-} phosphate H_3PO_4 phosphoric acid	ex. SO_4^{2-} sulfate H_2SO_4 sulfuric acid

Naming Molecular Compounds Study Guide

Gilleland, Gloria Mumford



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