

## AP Chemistry: Equilibrium Multiple Choice

57. For the reaction  $A_{(g)} \rightleftharpoons B_{(g)} + C_{(g)}$ , the equilibrium constant,  $K_p$ , is  $2 \times 10^{-4}$  at  $25^\circ\text{C}$ . A mixture of the three gases at  $25^\circ\text{C}$  is placed in a reaction flask and the initial pressures are  $P_A = 2$  atm,  $P_B = 0.5$  atm, and  $P_C = 1$  atm. At the instant of mixing, which of the following is true for the reaction as written?

- (A)  $\Delta G < 0$    (B)  $\Delta G > 0$    (C)  $\Delta S = 0$    (D)  $\Delta G^\circ = 0$    (E)  $\Delta G^\circ < 0$



Consider the equilibrium above. Which of the following changes will increase the concentration of  $\text{HgI}_4^{2-}$ ?

- (A) Increasing the concentration of  $\text{OH}^-$    (B) Adding 6 M  $\text{HNO}_3$   
(C) Increasing the mass of  $\text{HgO}$  present   (D) Increasing the temperature   (E) Adding a catalyst

29. In which of the following systems would the number of moles of the substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?

- (A)  $\text{CO}_{(g)} + \text{NO}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + 1/2 \text{N}_{2(g)}$    (B)  $\text{N}_{2(g)} + 3 \text{H}_{2(g)} \rightleftharpoons 2 \text{NH}_{3(g)}$   
(C)  $\text{N}_{2(g)} + 2 \text{O}_{2(g)} \rightleftharpoons 2 \text{NO}_{2(g)}$    (D)  $\text{N}_2\text{O}_{4(g)} \rightleftharpoons 2 \text{NO}_{2(g)}$   
(E)  $\text{NO}_{(g)} + \text{O}_{3(g)} \rightleftharpoons \text{NO}_{2(g)} + \text{O}_{2(g)}$



Some  $\text{PCl}_3$  and  $\text{Cl}_2$  are mixed in a container at  $200^\circ\text{C}$  and the system reaches equilibrium according to the equation above. Which of the following causes an increase in the number of moles of  $\text{PCl}_3$  present at equilibrium?

- I. Decreasing the volume of the container  
II. Raising the temperature  
III. Adding a mole of He gas at constant volume  
(A) I only   (B) II only   (C) I and III only   (D) II and III only   (E) I, II, and III



Equal numbers of moles of  $\text{HCl}$  and  $\text{O}_2$  in a closed system are allowed to reach equilibrium as represented by the equation above. Which of the following must be true at equilibrium?

- I.  $[\text{HCl}]$  must be less than  $[\text{Cl}_2]$ .  
II.  $[\text{O}_2]$  must be greater than  $[\text{HCl}]$ .  
III.  $[\text{Cl}_2]$  must equal  $[\text{H}_2\text{O}]$ .  
(A) I only   (B) II only   (C) I and III only   (D) II and III only   (E) I, II, and III

65. The solubility of  $\text{CuI}$  is  $2 \times 10^{-8}$  molar. What is the solubility product constant,  $K_{sp}$ , for  $\text{CuI}$ ?

- (A)  $1.4 \times 10^{-3}$    (B)  $2 \times 10^{-8}$    (C)  $4 \times 10^{-12}$    (D)  $2 \times 10^{-12}$    (E)  $8 \times 10^{-18}$

# General Chemistry And Solubility Equilibrium Multiple Choice

**DJ Losen**



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