

THE ULTIMATE GENERAL CHEMISTRY 1 STUDY GUIDE

❖ Matter and Measurement

- ✔ Matter: anything that has mass and occupies space
- ✔ Properties of matter: physical properties and chemical properties
- ✔ States of matter: solid, liquid, gas
- ✔ Measurements: SI Units, significant figures, dimensional analysis

❖ Atoms, Molecules, and Ions

- ✔ Atomic structure: protons, neutrons, and electrons
- ✔ Periodic table and trends: electronegativity, ionization energy, electron affinity
- ✔ Chemical bonding: ionic bonding, covalent bonding, metallic bonding
- ✔ Chemical formulas and equations: empirical and molecular formulas, balancing chemical equations

❖ Stoichiometry

- ✔ Limiting reactants and percent yield
- ✔ Mole concept and Avogadro's number
- ✔ Molarity and molality
- ✔ Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law

❖ Chemical Reactions and Equations

- ✔ Types of chemical reactions: synthesis, decomposition, single replacement, double replacement, combustion
- ✔ Redox reactions and balancing redox equations
- ✔ Reaction rates: factors affecting reaction rates, rate laws, and reaction mechanisms

❖ Gases

- ✔ Gas laws: Boyle's law, Charles's law, Gay-Lussac's law, Avogadro's law, Ideal gas law
- ✔ Kinetic Molecular Theory of gases: pressure, temperature, volume, and the number of moles of gas particles
- ✔ Natural gases: van der Waals equation

General Chemistry Thinking Guide

Rosina Ehmann



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