

Isotopes Practice Set

- What does "atomic number" mean?
- What does the number next to isotopes mean?
- How do you tell isotopes apart?

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Magnesium-24	Magnesium-26
# of protons	12	12
# of neutrons	12	14
# of electrons	12	12

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Sulfur-32	Sulfur-34
# of protons	16	16
# of neutrons	16	18
# of electrons	16	16

Fill in the isotopes names including atomic numbers, including atomic numbers from the chart. Use your periodic table and the information provided.

	Isotope: $^{12}_6\text{C}$	Isotope: $^{16}_8\text{O}$
# of protons	6	8
# of neutrons	6	8
# of electrons	6	8

	Isotope: $^{14}_7\text{N}$	Isotope: $^{23}_{11}\text{Na}$
# of protons	7	11
# of neutrons	7	12
# of electrons	7	11

	Isotope: $^{14}_7\text{N}$	Isotope: $^{16}_8\text{O}$
# of protons	7	8
# of neutrons	7	8
# of electrons	7	8

	Isotope: $^{12}_6\text{C}$	Isotope: $^{13}_6\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

Isotopes Practice Set Key

Anthony S. Fauci



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