

Electron Configuration Practice Worksheet

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 5s^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4 6d^1$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^1 4d^{10}$ → $[Kr] 5s^1 4d^{10}$ (exception → more stable)
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$ ← last electron in of shell

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ Sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ Rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ Antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ Osmium
- 15) $[Rn] 7s^2 5f^{11}$ Terbium

Determine which of the following electron configurations are not valid:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (3d not there)
- 17) $1s^2 2s^2 2p^6 3s^2 3d^5$ not valid (3d comes after 3p)
- 18) $[Ra] 7s^2 5f^6$ not valid [Ra] is not a noble gas
- 19) $[Kr] 5s^2 4d^{10} 5p^5$ valid (iodine)
- 20) $[Xe]$ not valid (noble gases cannot be abbreviated)

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Table of Contents Electron Configuration Challenge Worksheet

1. Understanding the eBook Electron Configuration Challenge Worksheet
 - The Rise of Digital Reading Electron Configuration Challenge Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Challenge Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Challenge Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Challenge Worksheet
 - Personalized Recommendations
 - Electron Configuration Challenge Worksheet User Reviews and Ratings

- Electron Configuration Challenge Worksheet and Bestseller Lists
- 5. Accessing Electron Configuration Challenge Worksheet Free and Paid eBooks
 - Electron Configuration Challenge Worksheet Public Domain eBooks
 - Electron Configuration Challenge Worksheet eBook Subscription Services
 - Electron Configuration Challenge Worksheet Budget-Friendly Options
- 6. Navigating Electron Configuration Challenge Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Challenge Worksheet Compatibility with Devices
 - Electron Configuration Challenge Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Challenge Worksheet
 - Highlighting and Note-Taking Electron Configuration Challenge Worksheet
 - Interactive Elements Electron Configuration Challenge Worksheet
- 8. Staying Engaged with Electron Configuration Challenge Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Challenge Worksheet
- 9. Balancing eBooks and Physical Books Electron Configuration Challenge Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Challenge Worksheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Challenge Worksheet
 - Setting Reading Goals Electron Configuration Challenge Worksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Challenge Worksheet
 - Fact-Checking eBook Content of Electron Configuration Challenge Worksheet
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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