

ANSWER KEY

ATOMS AND ELECTRONS STUDY GUIDE

Learning Goals: (1) (10-1) P. 6-8 - Explain a new scientific theory of atomic-scale phenomena (atomic theory) by describing changes in the atomic model over time, including those changes that are associated with quantum mechanics. (10-1) P. 6-8 - Explain the relative physical sizes of atoms, nuclei, and electrons, showing how the relative sizes of particles determine interactions, including how the spacing between particles is related to their mass, electric charges, and magnetic spins, the spin. (10-1) P. 6-8 - Explain properties of atoms and their nucleus, such as isotopes, nuclear forces, composition of their nucleus, related to P. 6-8, and distinguish the theoretical description of their fundamental (strong, electromagnetic, weak, gravity, and spin) forces.

ANSWERS

Question: Match the atomic theory described in the correct conditions.

Model	Scientist(s)	Atomic Structure
(A)	Cathode ray (1897)	(A) Everything is made of atoms. Atoms cannot be created, divided or destroyed.
(B)	Bohr (1913)	(B) Matter is made of small particles called atoms. Atoms have a central nucleus with positive charge and electrons with negative charge are present around the nucleus.
(C)	Rutherford (1911)	(C) Atoms are indivisible particles.
(D)	Rutherford (1911)	(D) Group number = Valency element
(E)	Moseley (1913)	(E) Atomic mass
(F)	Electron (1897)	(F) Matter is made of very small particles called atoms. Atoms cannot be created or destroyed.

Question: Use the periodic table to answer the following questions:

- (1) **(10-1) P. 6-8** - The number of protons equals the number of electrons.
- (2) **(10-1) P. 6-8** - The atomic number will equal the **number of protons**.
- (3) **(10-1) P. 6-8** - **(10-1) P. 6-8** - The number of electrons in a neutral atom is equal to the number of protons. The number of electrons in an ion is equal to the number of protons minus the charge.

Electrons In Atoms Answer Key Study Guide

RD Boyd



Electrons In Atoms Answer Key Study Guide:

This book delves into Electrons In Atoms Answer Key Study Guide. Electrons In Atoms Answer Key Study Guide is a crucial topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Electrons In Atoms Answer Key Study Guide, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Electrons In Atoms Answer Key Study Guide
 - Chapter 2: Essential Elements of Electrons In Atoms Answer Key Study Guide
 - Chapter 3: Electrons In Atoms Answer Key Study Guide in Everyday Life
 - Chapter 4: Electrons In Atoms Answer Key Study Guide in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, this book will provide an overview of Electrons In Atoms Answer Key Study Guide. The first chapter will explore what Electrons In Atoms Answer Key Study Guide is, why Electrons In Atoms Answer Key Study Guide is vital, and how to effectively learn about Electrons In Atoms Answer Key Study Guide.
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 4. In chapter 3, the author will examine the practical applications of Electrons In Atoms Answer Key Study Guide in daily life. This chapter will showcase real-world examples of how Electrons In Atoms Answer Key Study Guide can be effectively utilized in everyday scenarios.
 5. In chapter 4, the author will scrutinize the relevance of Electrons In Atoms Answer Key Study Guide in specific contexts. This chapter will explore how Electrons In Atoms Answer Key Study Guide is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Electrons In Atoms Answer Key Study Guide. The final chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Electrons In Atoms Answer Key Study Guide.

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