

Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

## Student Exploration: Equilibrium and Concentration

**Essential Question:** How do changes in concentration affect chemical equilibrium?

**What You Will Do:** Explore how changes in concentration affect chemical equilibrium.

Many real-world systems are in chemical equilibrium. For example, the chemical reaction between nitrogen and hydrogen to form ammonia is an equilibrium reaction.

1. How does the concentration of nitrogen affect the equilibrium? **It shifts to the right.**

2. How does the concentration of hydrogen affect the equilibrium? **It shifts to the right.**

3. What happens if the concentration of ammonia is increased? **It shifts to the left.**

### What If...?

4. How does the concentration of nitrogen affect the equilibrium? **It shifts to the right.**

5. How does the concentration of hydrogen affect the equilibrium? **It shifts to the right.**

6. How does the concentration of ammonia affect the equilibrium? **It shifts to the left.**

**When the concentration of a reactant is increased, the equilibrium shifts to the right to produce more products. When the concentration of a product is increased, the equilibrium shifts to the left to produce more reactants.**

7. How does the concentration of nitrogen affect the equilibrium? **It shifts to the right.**

**When the concentration of a reactant is increased, the equilibrium shifts to the right to produce more products.**

8. How does the concentration of hydrogen affect the equilibrium? **It shifts to the right.**

**When the concentration of a reactant is increased, the equilibrium shifts to the right to produce more products.**



# Explorelearning Equilibrium And Concentration Gizmo Answer Key

**TD Snyder**



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