

Element Builder

Vocabulary: atom, atomic number, electron, electron dot diagram, element, energy level, ion, isotope, mass number, neutron, nucleus, periodic table, proton, radioactive, valence electrons

An **isotope** is an alternative form of an element. Each isotope of an element has the same number of protons, but a different number of neutrons. The isotope is represented by the atomic symbol and mass number. Some isotopes are stable, while others are **radioactive**, which means the atoms decay over time and emit radiation.

- What are the stable isotopes of carbon? _____
- What are the stable isotopes of nitrogen? _____
- List two radioactive isotopes of oxygen: _____

Use the Gizmo to answer the following questions.

- How many electrons are in a neutral atom of lithium? _____
- How many neutrons are in an atom of Mg-25? _____
- What is the mass number of an atom with 5 protons and 7 neutrons? _____
- An **ion** is a charged atom. How many electrons are in O^{2-} ? _____
- How many electrons are in Mg^{2+} ? _____

Question: How are electrons arranged around the nucleus of an atom?

Add electrons to your atom until you have used all the available electrons. How are the electrons arranged?

Electrons are arranged in orbits called **energy levels**, shown in the Gizmo.

- How many electrons can fit in the first energy level? _____
- How many electrons can fit in the second energy level? _____
- How many electrons can fit in the third energy level? _____

The electrons in the outermost orbit, called **valence electrons**, help to create chemical bonds. Create a lithium atom (3 protons, 4 neutrons, 3 electrons).

How many valence electrons are in a neutral lithium atom? _____

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