

# Key Terms

## Chemical Bonds

**Directions:** Match each term in Column I with its description in Column II. Write the letter of the correct term in the space provided.

### Column I

- \_\_\_\_\_ 1. binary compound
- \_\_\_\_\_ 2. chemically stable
- \_\_\_\_\_ 3. nonpolar molecule
- \_\_\_\_\_ 4. ion
- \_\_\_\_\_ 5. ionic bond
- \_\_\_\_\_ 6. polar molecule
- \_\_\_\_\_ 7. oxidation number
- \_\_\_\_\_ 8. chemical formula
- \_\_\_\_\_ 9. covalent bond
- \_\_\_\_\_ 10. hydrate
- \_\_\_\_\_ 11. chemical bond
- \_\_\_\_\_ 12. polyatomic ion

### Column II

- a. number that indicates how many electrons an atom must gain, lose, or share to become stable
- b. shorthand that tells what elements a compound contains and the exact number of atoms of each element in a unit of the compound
- c. positively or negatively charged, covalently bonded group of atoms
- d. compound composed of two elements
- e. describes an atom that has a full outermost energy level
- f. molecule that has a slightly positive end and a slightly negative end
- g. the attraction that forms between atoms when they share electrons
- h. the force that holds atoms together in a compound
- i. a compound that has water chemically attached to it
- j. the force of attraction between the opposite charges of the ions in an ionic compound
- k. molecule made of two identical atoms that share the electrons equally
- l. a charged particle that has either more or fewer electrons than protons

# Glencoe Physical Science Test Answers

**Raffaella Di Napoli**



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