

Chapter 15 Alcohols Practice questions

Question number	Answer	Marks	Guidance						
1 (a)	A mixture of liquids is heated to boiling point for a prolonged time	1							
	Vapour is formed which escapes from the liquid mixture, is changed back into liquid and returned to the liquid mixture	1							
	Any ethanal and ethanol that initially evaporates can then be oxidised	1							
1 (b)	$\text{CH}_3\text{CH}_2\text{OH} + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{COOH} + 4\text{H}^+ + 4\text{e}^-$	1							
2 (a) (i)	Compounds with the same molecular formula	1	This is a definition so should be learnt.						
	But different structural formulae / different structures	1							
2 (a) (ii)	$\text{C}_2\text{H}_6\text{O}$ only	1							
2 (a) (iii)	CH_2 only	1							
2 (b)	Potassium dichromate(VI) / $\text{K}_2\text{Cr}_2\text{O}_7$ and acid / acidified / H_2SO_4 / H^+	1	You can also have KMnO_4 / H_2SO_4 , but not HCl . Or remain purple if KMnO_4 used. Or goes from purple to colourless if KMnO_4 in acid used or gives brown precipitate or goes green if KMnO_4 neutral or in alkali.						
	Remains orange or no change or no reaction	1							
	Orange to green	1							
2 (c)	<u>Choice of reagents</u>	1	These tests always come up, so learn the tests to distinguish aldehydes and ketones. If you can't learn them all pick one to really learn.						
	<table border="1" style="width: 100%;"> <tr> <td style="width: 33%;">potassium dichromate (VI) / $\text{K}_2\text{Cr}_2\text{O}_7$ and acid / acidified / H_2SO_4 / H^+ or KMnO_4 / H_2SO_4</td> <td style="width: 33%;">Fehling's / Benedict's reagent</td> <td style="width: 33%;">Tollens' reagent or AgNO_3 / NH_3 or ammoniacal silver nitrate Not AgNO_3 alone.</td> </tr> </table>			potassium dichromate (VI) / $\text{K}_2\text{Cr}_2\text{O}_7$ and acid / acidified / H_2SO_4 / H^+ or KMnO_4 / H_2SO_4	Fehling's / Benedict's reagent	Tollens' reagent or AgNO_3 / NH_3 or ammoniacal silver nitrate Not AgNO_3 alone.			
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	with the aldehyde C			1					
<table border="1" style="width: 100%;"> <tr> <td style="width: 40%;">goes orange to green</td> <td style="width: 20%;">red solid</td> <td style="width: 40%;">silver mirror</td> </tr> <tr> <td>goes purple to colourless / brown ppt. / green solution if KMnO_4 used red solid</td> <td></td> <td></td> </tr> <tr> <td>Potassium dichromate is more usual to do in school.</td> <td></td> <td></td> </tr> </table>	goes orange to green	red solid	silver mirror		goes purple to colourless / brown ppt. / green solution if KMnO_4 used red solid			Potassium dichromate is more usual to do in school.	
goes orange to green	red solid	silver mirror							
goes purple to colourless / brown ppt. / green solution if KMnO_4 used red solid									
Potassium dichromate is more usual to do in school.									
with the ketone D	1								
<table border="1" style="width: 100%;"> <tr> <td style="width: 33%;">remains orange or no change or no reaction or purple for KMnO_4</td> <td style="width: 33%;">remains blue or no change or no reaction</td> <td style="width: 33%;">remains colourless or no change or no reaction</td> </tr> </table>		remains orange or no change or no reaction or purple for KMnO_4	remains blue or no change or no reaction	remains colourless or no change or no reaction					
remains orange or no change or no reaction or purple for KMnO_4	remains blue or no change or no reaction	remains colourless or no change or no reaction							

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