

Honors Chemistry Final Exam Review- 2021

- Make sure you have the following: scientific calculator and a pencil or pen with blue/black ink.
- The final exam will count as a double summative grade.
- The final exam is based on content from 2nd semester.

Moles

- How many particles are in a mole? 6.02×10^{23}
- What is the conversion between moles and mass? $\text{moles} = \frac{\text{molar mass}}$
- What is the conversion between moles and number of particles? $\text{moles} = \frac{6.02 \times 10^{23}}$
- What is the conversion between moles and volume of gas at STP? $\text{moles} = \frac{22.4 \text{ L}}$
- Do the following conversions:
 - How many grams are in 0.354 mol of cholesterol, $\text{C}_{27}\text{H}_{46}\text{O}$? 137 g
 - How many formula units are in 12.3g potassium nitrate? How many atoms? KNO_3 7.32×10^{22} Form U.
 - What volume at STP would 3.2g of carbon dioxide occupy? CO_2 1.6 L
 - How many molecules of helium are in 2.0L balloon at STP? He 5.4×10^{22} molec.
 - How many ~~atoms~~ are present in 3.2L of carbon dioxide at STP? CO_2 8.6×10^{22} atoms
- Do the following percent composition problems:
 - What is the percent of carbon in carbon dioxide? 27.29%
 - In a sample of pitchblende, uranium ore, it is determined that there are 2.5 g of uranium in a 500.0 g sample. What percent of the ore is uranium? 0.5%
 - What mass of carbon can you expect to recover from a 5.00g sample of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$)? How many carbon atoms are present? 12.0 g C 1.00×10^{23} atoms
 - A piece of jewelry is composed of 0.977mol of metal atoms. If 80.0% of the atoms are silver, how many silver atoms are in the jewelry? 0.782 mol 4.71×10^{23} atoms
- Work the following molecular/empirical formula problems:
 - What is the empirical formula for a compound if 4.00g is zinc and 1.00g is oxygen? ZnO
 - A gaseous compound has a percentage composition of 82.4% nitrogen and 17.6 % hydrogen. What is the empirical formula of the gas? NH_3
 - What is the molecular formula for a compound with an empirical formula of NO_2 and a molecular mass of 92.0g/mol? N_2O_4
 - What is the empirical formula for a compound if a 32.8g sample contains 22.8g of barium, 2.00g of carbon, and 8.00 g of oxygen? BaCO_3
 - Identify which of the following are empirical formulas. H_2O C_6H_{12} H_2O_2 C_2H_4 Na_2PO_4
 - A 50.0g sample of sodium carbonate hydrate is heated. After heating, 3.71g of the anhydrous salt remains. What is the formula for this hydrate? Name this hydrate. $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$
Sodium carbonate decahydrate

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