

- Here are three isotopes of an element: $^{12}_6\text{C}$ $^{13}_6\text{C}$ $^{14}_6\text{C}$
 - The element is: **Carbon**
 - The number 6 refers to the: **Atomic Number**
 - The numbers 12, 13, and 14 refer to the **different mass numbers**
 - How many protons and neutrons are in the first isotope? **P: 6 N: 6**
 - How many protons and neutrons are in the second isotope? **P: 6 N: 7**
 - How many protons and neutrons are in the third isotope? **P: 6 N: 8**

2. Complete the following chart:

Isotope name	atomic #	mass #	# of protons	# of neutrons	# of electrons
Potassium-37	19	37	19	18	19
Oxygen-17	8	17	8	9	8
uranium-235	92	235	92	143	92
uranium-238	92	238	92	146	92
boron-10	5	10	5	5	5
boron-11	5	11	5	6	5

3. How do you calculate the mass number of an element?

Mass # = protons + neutrons

- If an atom has 14 protons, 17 neutrons and 14 electrons, which expression identifies the element?
 - Silicon-14
 - Silicon-28
 - Silicon-31**
 - Silicon-17
- If an atom has 53 protons, 54 neutrons and 53 electrons, which expression identifies the element?
 - Iodine-53
 - Iodine-127
 - Iodine-54
 - Iodine-107**

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Table of Contents Isotopes Practice Set Answer Key

1. Understanding the eBook Isotopes Practice Set Answer Key
 - The Rise of Digital Reading Isotopes Practice Set Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Isotopes Practice Set Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Isotopes Practice Set Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Isotopes Practice Set Answer Key

- Personalized Recommendations
 - Isotopes Practice Set Answer Key User Reviews and Ratings
 - Isotopes Practice Set Answer Key and Bestseller Lists
5. Accessing Isotopes Practice Set Answer Key Free and Paid eBooks
 - Isotopes Practice Set Answer Key Public Domain eBooks
 - Isotopes Practice Set Answer Key eBook Subscription Services
 - Isotopes Practice Set Answer Key Budget-Friendly Options
 6. Navigating Isotopes Practice Set Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Isotopes Practice Set Answer Key Compatibility with Devices
 - Isotopes Practice Set Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Isotopes Practice Set Answer Key
 - Highlighting and Note-Taking Isotopes Practice Set Answer Key
 - Interactive Elements Isotopes Practice Set Answer Key
 8. Staying Engaged with Isotopes Practice Set Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Isotopes Practice Set Answer Key
 9. Balancing eBooks and Physical Books Isotopes Practice Set Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Isotopes Practice Set Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Isotopes Practice Set Answer Key
 - Setting Reading Goals Isotopes Practice Set Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Isotopes Practice Set Answer Key

- Fact-Checking eBook Content of Isotopes Practice Set Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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