

Mol = mass/Mr

Mol = concentration (mol/dm³) x volume (dm³)

Volume (dm³) = mol/24

1 cm³ = 0.001 dm³ (divided by 1,000)

What is a mol? One mole of a substance is equal to 6.022×10^{23} units of that substance (such as atoms, molecules, or ions). The number 6.022×10^{23} is known as Avogadro's constant.

Avogadro constant = 6.022×10 to the power of 23

Percentage mass:

Percentage mass = the percentage by mass of an element in the compound

E.g % by mass of Fe in Fe₂O₃ = $100 \times 2(56)/160 = 70\%$

Moles:

E.g what mass of oxygen is reacted when 6g of magnesium reacts?



Step 1: Use the equation mol = mass/Mr to find the number of moles of magnesium that reacted

$$6\text{g}/24\text{g} = 0.25 \text{ mol}$$

Step 2: Find the ratio between magnesium and oxygen in the reaction

$$\text{Mg}:\text{O}_2 = 2:1$$

$$0.25/2 = 0.125$$

Step 3: Calculate the mass of oxygen that was reacted

$$\text{Mass} = \text{mol} \times \text{Mr}$$

$$\text{Mr of O}_2 = 32$$

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insert a visualization in the device tree then the visualization manager is inserted automatically with the `targetvisu` and `webvisu` visualization types in addition another visu task is