

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the the systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
4. In part 1, what is the conjugate acid and base? Write the K_{eq} for the reaction. Is the reaction endothermic or exothermic? Explain the effect of temperature on the reaction. How does the color change? Write the balanced equation for the reaction in part 1.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
5. Write the balanced equation for the reaction in part 4.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
6. Write the balanced equation for the reaction in part 5.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
7. Write the balanced equation for the reaction in part 6.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
8. Write the balanced equation for the reaction in part 7.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
9. Write the balanced equation for the reaction in part 8.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
10. Write the balanced equation for the reaction in part 9.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$

Investigating Chemical Equilibrium Lab 12a

Marco Casella



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Investigating Chemical Equilibrium Lab 12a Book Review: Unveiling the Power of Words

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Table of Contents Investigating Chemical Equilibrium Lab 12a

1. Understanding the eBook Investigating Chemical Equilibrium Lab 12a
 - The Rise of Digital Reading Investigating Chemical Equilibrium Lab 12a
 - Advantages of eBooks Over Traditional Books
2. Identifying Investigating Chemical Equilibrium Lab 12a
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Investigating Chemical Equilibrium Lab 12a
 - User-Friendly Interface
4. Exploring eBook Recommendations from Investigating Chemical Equilibrium Lab 12a
 - Personalized Recommendations
 - Investigating Chemical Equilibrium Lab 12a User Reviews and Ratings
 - Investigating Chemical Equilibrium Lab 12a and Bestseller Lists

5. Accessing Investigating Chemical Equilibrium Lab 12a Free and Paid eBooks
 - Investigating Chemical Equilibrium Lab 12a Public Domain eBooks
 - Investigating Chemical Equilibrium Lab 12a eBook Subscription Services
 - Investigating Chemical Equilibrium Lab 12a Budget-Friendly Options
6. Navigating Investigating Chemical Equilibrium Lab 12a eBook Formats
 - ePub, PDF, MOBI, and More
 - Investigating Chemical Equilibrium Lab 12a Compatibility with Devices
 - Investigating Chemical Equilibrium Lab 12a Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Investigating Chemical Equilibrium Lab 12a
 - Highlighting and Note-Taking Investigating Chemical Equilibrium Lab 12a
 - Interactive Elements Investigating Chemical Equilibrium Lab 12a
8. Staying Engaged with Investigating Chemical Equilibrium Lab 12a
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Investigating Chemical Equilibrium Lab 12a
9. Balancing eBooks and Physical Books Investigating Chemical Equilibrium Lab 12a
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Investigating Chemical Equilibrium Lab 12a
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Investigating Chemical Equilibrium Lab 12a
 - Setting Reading Goals Investigating Chemical Equilibrium Lab 12a
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Investigating Chemical Equilibrium Lab 12a
 - Fact-Checking eBook Content of Investigating Chemical Equilibrium Lab 12a
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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