

Gas Variables

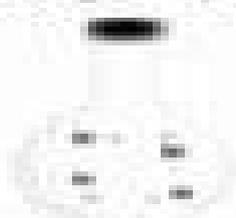
How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and amount (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 2 units

External pressure = 1 atm Internal pressure = 1 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



A2

Volume = 2 units

External pressure = 1 atm Internal pressure = 2 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K

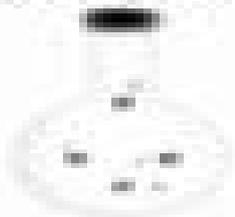


A3

Volume = 2 units

External pressure = 1 atm Internal pressure = 3 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K

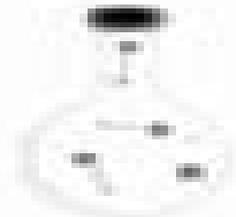
Experiment B (Heating the gas)



B1

Volume = 2 units

External pressure = 1 atm Internal pressure = 1 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



B2

Volume = 2 units

External pressure = 1 atm Internal pressure = 2 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 400 K Temperature = 400 K Temperature = 400 K



B3

Volume = 2 units

External pressure = 1 atm Internal pressure = 3 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 600 K Temperature = 600 K Temperature = 600 K

Note: Volume is the total number of units rather than three because 3 molecules of gas at these conditions gain sufficient kinetic energy each second ($v = 300\text{ m/s}$) that they collide themselves as much as they collide with the walls of the container.

Gas Variables Packet Answer Key

Michael Seilmaier



Gas Variables Packet Answer Key:

The book delves into Gas Variables Packet Answer Key. Gas Variables Packet Answer Key is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Gas Variables Packet Answer Key, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Gas Variables Packet Answer Key
- Chapter 2: Essential Elements of Gas Variables Packet Answer Key
- Chapter 3: Gas Variables Packet Answer Key in Everyday Life
- Chapter 4: Gas Variables Packet Answer Key in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Gas Variables Packet Answer Key. The first chapter will explore what Gas Variables Packet Answer Key is, why Gas Variables Packet Answer Key is vital, and how to effectively learn about Gas Variables Packet Answer Key.
3. In chapter 2, this book will delve into the foundational concepts of Gas Variables Packet Answer Key. The second chapter will elucidate the essential principles that need to be understood to grasp Gas Variables Packet Answer Key in its entirety.
4. In chapter 3, this book will examine the practical applications of Gas Variables Packet Answer Key in daily life. This chapter will showcase real-world examples of how Gas Variables Packet Answer Key can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Gas Variables Packet Answer Key in specific contexts. The fourth chapter will explore how Gas Variables Packet Answer Key is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Gas Variables Packet Answer Key. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Gas Variables Packet Answer Key.

https://blog.gospelcrusade.org/book/virtual-library/Documents/In_The_Analogy.pdf

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