

Ideal Gas Law Packet

Name _____

12.3

Date _____ Period _____

Given: Ideal Gas Law =

then P =

n =

V =

T =

R =

1. What pressure is required to contain 0.023 moles of nitrogen gas in a 4.2 L container at a temperature of 30. °C?
2. Oxygen gas is collected at a pressure of 123 kPa in a container which has a volume of 10.0 L. What temperature must be maintained on 0.500 moles of this gas in order to maintain this pressure? Express the temperature in degrees Celsius.
3. How many moles of chlorine gas would occupy a volume of 35.5 L at a pressure of 100.0 kPa and a temperature of 100. °C? After determining the number of moles, calculate the number of grams of chlorine (Cl₂) contained in this container?
4. What is the volume of a balloon if it contains 3.2 moles of helium at a temperature of 20. °C and standard pressure?
5. Calculate the volume which 1.00 mole of a gas occupies at STP.

Gas Law Packet

Rachel S Tattersall



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