

IDEAL GAS LAW

Name: _____

Use the Ideal Gas Law below to solve the following problems.

$PV = nRT$ where P is pressure in atmospheres,
 V is volume in liters,
 n is number of moles of gas,
 R is Ideal Gas Constant
 $0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$,
 T is Kelvin temperature.

moles of gas = n
 volume = V
 pressure = P

1. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is heated to 327 K. What is the new volume?

2. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is cooled to 127 K. What is the new volume?

3. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is heated to 327 K. What is the new pressure?

4. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is cooled to 127 K. What is the new pressure?

5. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is heated to 327 K. What is the new number of moles?

6. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is cooled to 127 K. What is the new number of moles?

7. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is heated to 327 K. What is the new density?

8. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is cooled to 127 K. What is the new density?

9. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is heated to 327 K. What is the new molar mass?

10. A gas sample with a volume of 1.00 liter at a pressure of 1.00 atm and a temperature of 273 K is cooled to 127 K. What is the new molar mass?

moles of gas = n

volume = V

pressure = P

moles of gas = n

volume = V

$PV = nRT$
 $1.00 \text{ L} \cdot \text{atm} = n \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 273 \text{ K}$
 $n = 4.46 \text{ mol}$

$PV = nRT$
 $1.00 \text{ L} \cdot \text{atm} = n \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 327 \text{ K}$
 $n = 3.74 \text{ mol}$

$PV = nRT$
 $1.00 \text{ L} \cdot \text{atm} = 4.46 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot T$
 $T = 348 \text{ K}$

$PV = nRT$
 $1.00 \text{ L} \cdot \text{atm} = 3.74 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot T$
 $T = 327 \text{ K}$

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