

# Gas Laws Problem Packet

**Gas Laws Packet**

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

**Essentials**

Standard Constant:

Unit	Standard Pressure
atm	1.013
mmHg	760
cmHg	76
torr	760
atm	1.013

**22.4 L = 1 mol at STP**

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law:	$P_1V_1 = P_2V_2$
Charles's Law:	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law:	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law:	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law:	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Graham's Law:	$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- 1 atm,  $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$
- 1 atm,  $R = 8.314 \frac{\text{J}}{\text{mol}\cdot\text{K}}$

**Always Remember:**

- Temperature must always be in Kelvin! ( $K = ^\circ\text{C} + 273$ )
- If you are given two different units of pressure, CONVERT!
- Circle two columns before solving, initial and final, write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you're given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (MAYBE)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things!

**Charles' Law  $V \propto T$**

Directions: Complete each of the problems below. Show how you set up the problem and solve.

- The temperature inside my refrigerator is about 4°C. If I place a balloon in the refrigerator, will the volume of air in the balloon increase or decrease? Explain your answer. (15 pts)
- A balloon is inflated to a volume of 2.0 L at 27°C. If the temperature is lowered to 0°C, what will the volume of the balloon be? (15 pts)

**Gay-Lussac's Law  $P \propto T$**

- A gas is held at a constant volume. The pressure is 1.0 atm at 27°C. If the temperature is raised to 127°C, what will the pressure be? (15 pts)
- A gas is held at a constant volume. The pressure is 1.0 atm at 27°C. If the temperature is lowered to 0°C, what will the pressure be? (15 pts)

**Boyle's Law  $P \propto \frac{1}{V}$**

- A gas is held at a constant temperature. The pressure is 1.0 atm and the volume is 2.0 L. If the pressure is increased to 2.0 atm, what will the volume be? (15 pts)
- A gas is held at a constant temperature. The pressure is 1.0 atm and the volume is 2.0 L. If the pressure is decreased to 0.5 atm, what will the volume be? (15 pts)

**Avogadro's Law  $V \propto n$**

- A gas is held at a constant temperature and pressure. The volume is 2.0 L and there are 1.0 mol of gas. If there are 2.0 mol of gas, what will the volume be? (15 pts)
- A gas is held at a constant temperature and pressure. The volume is 2.0 L and there are 1.0 mol of gas. If there are 0.5 mol of gas, what will the volume be? (15 pts)

**Combined Gas Law  $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$**

- A gas is held at a constant temperature and pressure. The volume is 2.0 L, the pressure is 1.0 atm, and the temperature is 27°C. If the pressure is increased to 2.0 atm and the temperature is raised to 127°C, what will the volume be? (15 pts)
- A gas is held at a constant temperature and pressure. The volume is 2.0 L, the pressure is 1.0 atm, and the temperature is 27°C. If the pressure is decreased to 0.5 atm and the temperature is lowered to 0°C, what will the volume be? (15 pts)

**Graham's Law  $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$**

- Gas A has a molecular mass of 16 g/mol and a rate of effusion of 1.0. Gas B has a molecular mass of 64 g/mol. What is the rate of effusion of Gas B? (15 pts)
- Gas A has a molecular mass of 16 g/mol and a rate of effusion of 1.0. Gas B has a molecular mass of 64 g/mol. What is the rate of effusion of Gas A? (15 pts)

All Gas Laws Included!



# Gas Law Problems Packet

**VM Jensen**



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## **Table of Contents Gas Law Problems Packet**

1. Understanding the eBook Gas Law Problems Packet
  - The Rise of Digital Reading Gas Law Problems Packet
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Problems Packet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Law Problems Packet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Problems Packet
  - Personalized Recommendations
  - Gas Law Problems Packet User Reviews and Ratings
  - Gas Law Problems Packet and Bestseller Lists

5. Accessing Gas Law Problems Packet Free and Paid eBooks
  - Gas Law Problems Packet Public Domain eBooks
  - Gas Law Problems Packet eBook Subscription Services
  - Gas Law Problems Packet Budget-Friendly Options
6. Navigating Gas Law Problems Packet eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Law Problems Packet Compatibility with Devices
  - Gas Law Problems Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Law Problems Packet
  - Highlighting and Note-Taking Gas Law Problems Packet
  - Interactive Elements Gas Law Problems Packet
8. Staying Engaged with Gas Law Problems Packet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Law Problems Packet
9. Balancing eBooks and Physical Books Gas Law Problems Packet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Law Problems Packet
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Law Problems Packet
  - Setting Reading Goals Gas Law Problems Packet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Law Problems Packet
  - Fact-Checking eBook Content of Gas Law Problems Packet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- 
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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