

More Practice with Gas Laws **KEY**

Directions: For each question, identify the applicable law and solve for the correct answer using dimensional analysis. Express your answer to the correct number of significant figures. Use scientific notation where appropriate. Record all work in your science notebooks.

I. Gas Law Problems

1. A sample of hydrogen gas occupies a volume of 1.8 L at 0.80 atm. What would its volume be at 1.0 atm?

Boyle's Law $P_1V_1 = P_2V_2$

$P_1 = 0.80 \text{ atm}$ $V_1 = 1.8 \text{ L}$

$P_2 = 1.0 \text{ atm}$ $V_2 = ??$

$(.8 \text{ atm})(1.8 \text{ L}) = (1.0 \text{ atm})(V_2)$

$1.44 \text{ atm} \times \text{L} = 1 \text{ atm } V_2$

$V_2 = 1.4 \text{ L}$

2. A balloon full of helium has a volume of 1.75 L at a temperature of 28°C. What would the balloon's volume be at 50°C?

Charles' Law

$\frac{V_1}{T_1} = \frac{V_2}{T_2}$

$V_1 = 1.75 \text{ L}$ $T_1 = 301 \text{ K}$

$V_2 = ??$ $T_2 = 323 \text{ K}$

$V_2 = 1.9 \text{ L}$

3. If you have 5.00 mol of fluorine gas at 273 K and 1.00 atm of pressure, what will be the volume of the gas?

Ideal Gas Law

$PV = nRT$

$(1.00 \text{ atm})(V) = 5.00 \text{ mol} (.0821 \text{ L} \times \text{atm/mol} \times \text{K})(273 \text{ K})$

$V = 112 \text{ L} = 1.12 \times 10^2 \text{ L}$

Gas Laws Practice Calculations Answers Unit 09 Lesson 01

Katrin Zwirglmaier



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Table of Contents Gas Laws Practice Calculations Answers Unit 09 Lesson 01

1. Understanding the eBook Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - The Rise of Digital Reading Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Practice Calculations Answers Unit 09 Lesson 01

- Personalized Recommendations
 - Gas Laws Practice Calculations Answers Unit 09 Lesson 01 User Reviews and Ratings
 - Gas Laws Practice Calculations Answers Unit 09 Lesson 01 and Bestseller Lists
5. Accessing Gas Laws Practice Calculations Answers Unit 09 Lesson 01 Free and Paid eBooks
- Gas Laws Practice Calculations Answers Unit 09 Lesson 01 Public Domain eBooks
 - Gas Laws Practice Calculations Answers Unit 09 Lesson 01 eBook Subscription Services
 - Gas Laws Practice Calculations Answers Unit 09 Lesson 01 Budget-Friendly Options
6. Navigating Gas Laws Practice Calculations Answers Unit 09 Lesson 01 eBook Formats
- ePub, PDF, MOBI, and More
 - Gas Laws Practice Calculations Answers Unit 09 Lesson 01 Compatibility with Devices
 - Gas Laws Practice Calculations Answers Unit 09 Lesson 01 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - Highlighting and Note-Taking Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - Interactive Elements Gas Laws Practice Calculations Answers Unit 09 Lesson 01
8. Staying Engaged with Gas Laws Practice Calculations Answers Unit 09 Lesson 01
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Practice Calculations Answers Unit 09 Lesson 01
9. Balancing eBooks and Physical Books Gas Laws Practice Calculations Answers Unit 09 Lesson 01
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Practice Calculations Answers Unit 09 Lesson 01
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws Practice Calculations Answers Unit 09 Lesson 01
- Setting Reading Goals Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws Practice Calculations Answers Unit 09 Lesson 01

- Fact-Checking eBook Content of Gas Laws Practice Calculations Answers Unit 09 Lesson 01
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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