

Lab #8: Exploring the Properties of Gases

Purpose

In this experiment you will explore three gas laws Boyle's Law, Charles's Law, and Gay-Lussac's Law also known as Amontons's Law. The first activity relates pressure and volume. The second activity relates temperature and volume. The third activity relates temperature and pressure. Data will be gathered for each gas law, plotted as graphs and the necessary calculations will be completed to evaluate the gas law in each activity.

Introduction

The purpose of this investigation is to conduct three activities to study the empirical gas laws, Boyle's Law, Amontons's Law, and Charles's Law. Three properties of gases will be investigated: pressure, volume, and temperature. You will be using an online simulation to change the one of these variables and examine the effect on the others. By gathering and analyzing your data and observations, you will be able to describe the gas laws, both qualitatively and mathematically.

Boyle's Law

According to Boyle's Law, the pressure (P) of a gas varies inversely with the volume (V) of the gas when the temperature (T) and moles of gas (n) are kept constant.

$$P \propto \frac{1}{V} \quad (T \text{ and } n \text{ are constant})$$

According to Boyle's Law, volume decreases when the pressure increases, and volume increases when the pressure decreases.

$$P_2 \times V_2 = P_1 \times V_1$$

In this lab you will start with a known volume of gas and a pressure sensor will measure the pressure of the gas. You will decrease the volume of gas in increments, each time measuring the new pressure of the gas.

Amontons's Law

According to Amontons's Law, the pressure of a gas varies proportionally with the Kelvin temperature as long as the volume and quantity of gas are kept constant.

Gases Properties And Laws Lab Answers

Rachel Sandford



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