

① Atoms



Particle	P	N	E
Charge	+1	0	-1
Mass	1	1	0

Isotopes - A different atomic form of the same element, has the same number of protons, different number of neutrons.

¹² C	¹³ C
¹⁴ C	¹⁵ C
¹⁶ C	¹⁷ C
¹⁸ C	¹⁹ C

Relative atomic mass (Ar) =
 Sum of (isotope abundance x isotope mass no.)

 Sum of abundances of all the isotopes

② Compounds

A substance made of atoms of at least two different elements, chemically bonded together.

③ Mixtures

A substance made from two or more elements or compounds that aren't chemically bonded together.

④ Elements

All atoms of an element have the same number of protons.

Different elements have atoms with different numbers of protons.

⑤ Chemical Equations

During a chemical reaction bonds are broken, atoms rearrange, new bonds form



Balance:



Topic 1a -
Atoms, Elements,
Compounds &
Mixtures

⑥ Electronic Configuration

Configuration

Electrons occupy shells (energy levels)
 Electron shells with lowest energy is filled first

Shell	Max no. of electrons
1	2
2	8
3	8

⑦ Separating Techniques

There are a number of physical techniques used to separate mixtures of solids, liquids & gases.

Technique	What is separated
Paper Chromatography	Mix of different liquids
Filtration	Insoluble solids from liquid
Evaporation & Crystallisation	Soluble solids from liquid
Distillation	Mixture of liquids depending on boiling point

⑧ The History of the Atom

Year	Scientist	Model	Structure
1803	John Dalton	Solid sphere	
1904	J. J. Thompson	Plum pudding	
1911	Ernest Rutherford	Nuclear	
1913	Niels Bohr	Planetary	

Gcse Science Chemistry Revision Guide

John Scottow



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