

# NUMERICALS (CHEM-I)



Dr. ZAIN UL ABDEEN

## CHAPTER # 2

## GRAHAM'S LAW OF DIFFUSION

QUESTION # 24, pg. 72, 5TB8

Helium gas takes 5 seconds to effuse from a hole of 10 dm<sup>3</sup> container, how long it would take for 10 dm<sup>3</sup> of oxygen gas to effuse from the same container at the same temperature and pressure?

DATA:

$$t_{He} = 5 \text{ s}$$

$$M_{He} = 4 \text{ amu}$$

$$M_{O_2} = 32 \text{ amu}$$

$$t_{O_2} = ?$$

$$\frac{t_A}{t_B} = \sqrt{\frac{M_A}{M_B}}$$

Direct Relation

Finding  $t_{O_2}$  :

Hint: (The substance which is required should taken as numerator)

Using Graham's Law

$$\frac{t_{O_2}}{t_{He}} = \sqrt{\frac{M_{O_2}}{M_{He}}}$$

$$\frac{t_{O_2}}{5} = \sqrt{\frac{32}{4}}$$

$$\frac{t_{O_2}}{t_{He}} = \sqrt{8}$$

$$\frac{t_{O_2}}{5} = 2.828$$

$$t_{O_2} = 2.828 \times 5$$

$$t_{O_2} = 14.14 \text{ s}$$

Ends...

# Grahams Law Of Diffusion Answer Key

**Xiaolong Qi**



## **Grahams Law Of Diffusion Answer Key:**

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### **Table of Contents Grahams Law Of Diffusion Answer Key**

1. Understanding the eBook Grahams Law Of Diffusion Answer Key
  - The Rise of Digital Reading Grahams Law Of Diffusion Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Grahams Law Of Diffusion Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grahams Law Of Diffusion Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grahams Law Of Diffusion Answer Key
  - Personalized Recommendations
  - Grahams Law Of Diffusion Answer Key User Reviews and Ratings

- Grahams Law Of Diffusion Answer Key and Bestseller Lists
- 5. Accessing Grahams Law Of Diffusion Answer Key Free and Paid eBooks
  - Grahams Law Of Diffusion Answer Key Public Domain eBooks
  - Grahams Law Of Diffusion Answer Key eBook Subscription Services
  - Grahams Law Of Diffusion Answer Key Budget-Friendly Options
- 6. Navigating Grahams Law Of Diffusion Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Grahams Law Of Diffusion Answer Key Compatibility with Devices
  - Grahams Law Of Diffusion Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grahams Law Of Diffusion Answer Key
  - Highlighting and Note-Taking Grahams Law Of Diffusion Answer Key
  - Interactive Elements Grahams Law Of Diffusion Answer Key
- 8. Staying Engaged with Grahams Law Of Diffusion Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grahams Law Of Diffusion Answer Key
- 9. Balancing eBooks and Physical Books Grahams Law Of Diffusion Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grahams Law Of Diffusion Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Grahams Law Of Diffusion Answer Key
  - Setting Reading Goals Grahams Law Of Diffusion Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grahams Law Of Diffusion Answer Key
  - Fact-Checking eBook Content of Grahams Law Of Diffusion Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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