

Answer Key

Chapter 11

Practice Problems



2 molecules $\text{H}_2\text{O}_2 \rightarrow$ 1 molecule O_2 + 2 molecules H_2O

2 moles $\text{H}_2\text{O}_2 \rightarrow$ 1 mole O_2 + 2 moles H_2O

68.04 g $\text{H}_2\text{O}_2 =$ 32.00 g O_2 + 36.04 g H_2O , as shown below.

68.04 g reactant = 68.04 g products

$$2 \frac{\cancel{\text{mol H}_2\text{O}_2}}{\cancel{\text{mol H}_2\text{O}_2}} \times \frac{34.02 \text{ g H}_2\text{O}_2}{1 \cancel{\text{mol H}_2\text{O}_2}} = 68.04 \text{ g H}_2\text{O}_2$$

$$1 \frac{\cancel{\text{mol O}_2}}{\cancel{\text{mol O}_2}} \times \frac{32.00 \text{ g O}_2}{1 \cancel{\text{mol O}_2}} = 32.00 \text{ g O}_2$$

$$2 \frac{\cancel{\text{mol H}_2\text{O}}}{\cancel{\text{mol H}_2\text{O}}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \cancel{\text{mol H}_2\text{O}}} = 36.04 \text{ g H}_2\text{O}$$



1 formula unit $\text{H}_2\text{CO}_3 \rightarrow$ 1 molecule H_2O + 1 molecule CO_2

1 mole $\text{H}_2\text{CO}_3 \rightarrow$ 1 mole H_2O + 1 mole CO_2

62.03 g $\text{H}_2\text{CO}_3 =$ 18.02 g H_2O + 44.01 g CO_2 , as shown below.

62.03 g reactant = 62.03 g products

$$1 \frac{\cancel{\text{mol H}_2\text{CO}_3}}{\cancel{\text{mol H}_2\text{CO}_3}} \times \frac{62.03 \text{ g H}_2\text{CO}_3}{1 \cancel{\text{mol H}_2\text{CO}_3}} = 62.03 \text{ g H}_2\text{CO}_3$$

$$1 \frac{\cancel{\text{mol H}_2\text{O}}}{\cancel{\text{mol H}_2\text{O}}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \cancel{\text{mol H}_2\text{O}}} = 18.02 \text{ g H}_2\text{O}$$

$$1 \frac{\cancel{\text{mol CO}_2}}{\cancel{\text{mol CO}_2}} \times \frac{44.01 \text{ g CO}_2}{1 \cancel{\text{mol CO}_2}} = 44.01 \text{ g CO}_2$$

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Camilla Rothe



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